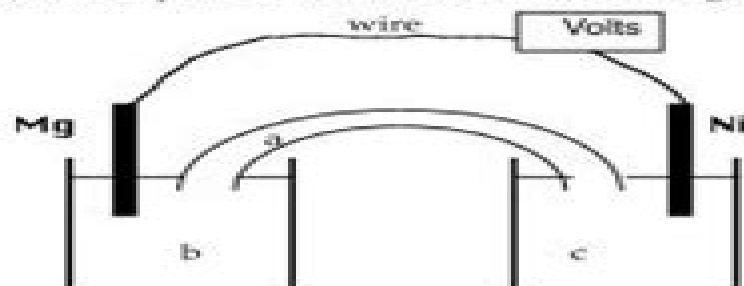


(Handout, p. 685-691, Notice: we calculate  $E^\circ_{\text{cell}}$  slightly differently than the book!)



1. Consider this voltaic cell made with Mg and Ni:

- part "a" is called the \_\_\_\_\_
- What solution is in container "b"?  
Mg(NO<sub>3</sub>)<sub>2</sub> or Ni(NO<sub>3</sub>)<sub>2</sub>
- What solution is in container "c"?  
Mg(NO<sub>3</sub>)<sub>2</sub> or Ni(NO<sub>3</sub>)<sub>2</sub>

d. Does the *oxidation* or *reduction* occur at the Mg electrode? Is Mg the *red* wire or *black* wire?

Write the equation for this half reaction \_\_\_\_\_

e. Does the *oxidation* or *reduction* occur at the Ni electrode? Is Ni the *red* wire or *black* wire?

Write the equation for this half reaction \_\_\_\_\_

f. Write the overall reaction for the cell \_\_\_\_\_

g. Which way do electrons flow in the wire? \_\_\_\_\_

h. Calculate  $E^\circ_{\text{cell}}$  for this cell.

2. a. For the Fe | Fe<sup>2+</sup> and Cu | Cu<sup>2+</sup> half cells : Which is oxidized? \_\_\_\_\_ Which is reduced? \_\_\_\_\_

Write the half reactions for the oxidation and the reduction and calculate  $E^\circ_{\text{cell}}$ .

b. For the Al | Al<sup>3+</sup> and Ag | Ag<sup>+</sup> half cells : Which is oxidized? \_\_\_\_\_ Which is reduced? \_\_\_\_\_

Write the half reactions for the oxidation and the reduction and calculate  $E^\circ_{\text{cell}}$ .

3. A voltaic cell is made with Palladium ( Pd | Pd<sup>2+</sup> ) = red wire and I<sub>2</sub> | I<sup>-</sup> = black wire. The measured  $E^\circ_{\text{cell}}$  is + 0.41 Volts. I<sub>2</sub> | I<sup>-</sup> is found on your  $E^\circ_{\text{red}}$  chart. Pd | Pd<sup>2+</sup> is not.

a. Is Pd | Pd<sup>2+</sup> the oxidation or the reduction? Is I<sub>2</sub> | I<sup>-</sup> the oxidation or the reduction?

b. Calculate  $E^\circ_{\text{red}}$  for the Pd<sup>2+</sup> + 2 e<sup>-</sup> —> Pd reaction using this data.

c. Determine  $E^\circ_{\text{ox}}$  for Pd —> Pd<sup>2+</sup> + 2 e<sup>-</sup> \_\_\_\_\_

(Ans. 1d. ox, black, e. red, red, f. Mg + Ni<sup>2+</sup> —> Mg<sup>2+</sup> + Ni, g. Mg to Ni, h. +2.12 V; 2a. Fe, Cu, +0.78 V, b. Al, Ag, +2.46 V, 3a. red, ox, b. +0.95 V, c. -0.95V; 4a. +0.34 V, b. -0.34 V)

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