

Chapter 4 Atomic Structure**Section 4.2 The Structure of an Atom**
(pages 108–112)

This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.

Reading Strategy (page 108)

Monitoring Your Understanding Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Know About Atoms	What I Would Like to Learn	What I Have Learned
Most students will know that atoms are the “building blocks” of matter, and some may know that atoms contain subatomic particles.	Based on the title of the section, students may say that they want to learn more about the structure of atoms.	

Properties of Subatomic Particles (pages 108–109)

1. What are three subatomic particles?
a. Protons b. Electrons c. Neutrons
2. Circle the letter that identifies a subatomic particle with a positive charge.
a. nucleus b. proton c. neutron d. electron
3. Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? A charged object did not deflect the paths of the particles.

Comparing Subatomic Particles (pages 109–110)

4. Circle the letters of properties that vary among subatomic particles.
a. color b. mass
 c. charge d. location in the atom
5. Circle the letter of the expression that accurately compares the masses of neutrons and protons.
a. mass of 1 neutron = mass of 1 proton
b. mass of 2000 neutrons = mass of 1 proton
c. mass of 1 electron = mass of 1 proton
d. mass of 1 neutron = mass of 1 electron

Chapter 4 Atomic Structure Answers

EW Minium

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The Mining Magazine,1927 **On the Application of the Quantum Theory to Atomic Structure** Niels Bohr,1924 Study Guide with Answers to Selected Problems H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

Study Guide and Solutions Manual to accompany Basic Concepts of Chemistry, 9th Edition Leo J. Malone,2011-12-07 This is the Study Guide and Solutions Manual to accompany Malone s Basic Concepts of Chemistry Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed Peter P. Berlow,1978

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