

Chemistry Study Guide Chapter 4 - Key

1. Define the following terms: (a) atom, (b) atomic number, (c) mass number, (d) isotopes (give 2 definitions), (e) atomic mass unit (amu), (f) nucleus, (g) spectroscopy, (h) ground state, (i) excited state, (j) atomic mass, (k) quantum leap, (l) photon.

- (a) **Atom:** the smallest particle of an element that still has the element's properties.
- (b) **Atomic number:** the number of protons in the nucleus of an atom.
- (c) **Mass number:** the number of protons plus the number of neutrons in an atom.
- (d) **Isotopes:** Atoms with the same atomic number but different mass numbers or atoms with the same number of protons but a different number of neutrons.
- (e) **Atomic mass unit (amu):** small unit of mass used to measure atoms, subatomic particles, etc.
- (f) **Nucleus:** small, dense, positively charged core of an atom containing protons and neutrons.
- (g) **Spectroscopy:** the study of spectra produced by the excitation of atoms, especially experimental observation of optical spectra used to determine a substance's identity.
- (h) **Ground state:** the state of lowest energy for an atom.
- (i) **Excited state:** any state of higher energy for an atom.
- (j) **Atomic mass:** the weighted average of all the isotopes of an element.
- (k) **Quantum Leap:** change of an electron from one energy state to another within an atom.
- (l) **Photon:** small packets of electromagnetic energy

2. Compare and contrast electrons, protons, and neutrons in terms of mass, charge and location.

PARTICLE	SYMBOL	CHARGE	MASS (amu)	LOCATION
electron	e^{-}	-1	≈ 0	orbit nucleus
proton	p^{+}	+1	≈ 1	inside nucleus
neutron	n^{0}	0	≈ 1	inside nucleus

3. What are the units used to express (a) wavelength, (b) frequency.

- (a) **any units of length** (b) **cycles/second or hertz.**

4. Compare/contrast x-rays and infrared radiation in terms of their speed, wavelength, frequency and energy.

Both move at the speed of light however x-rays have shorter wavelengths which give them higher frequencies and energies than infrared radiation.

Chapter 4 Chemistry Study Guide Answers

Susan McMurry



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