

### Phase Change

- Deposition-gas to a solid; Sublimation-solid to gas; Vaporization- liquid to gas; Condensation-gas to liquid
- Physical change - melting, freezing, dissolving, evaporating;
- Chemical change- boiling an egg, converting, exploding, burning, digesting.

### Atomic Numbers

- Number that identifies an element-Atomic number
- Weighted average mass of all occurring isotopes of an element-Atomic mass
- Symbolic method for expressing the composition of an atomic nucleus-Atomic notation

### Electron Configurations

- Shorthand description of electrons by sublevels according to increasing energy-Electron configuration
- Sublevels: 2,6,10,14 electrons in the sublevels; 1s 2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d 7p;



### Chemical Formulas

- Name of a chemical compound that indicates the # of atoms of each element-Chemical formula
- Number of atoms from formula- $\text{C}_7\text{H}_{16}\text{NO}_3$  = 40 atoms

$\text{NaHCO}_3$ -sodium hydrogen carbonate;  $\text{Mg}(\text{ClO}_4)_2$  - Magnesium Perchlorate;  $\text{PbO}_2$  - Lead (IV) oxide;  
 $\text{Hg}_2\text{Br}_2$ -Mercury (II) bromide;  $\text{Al}_2(\text{CrO}_4)_3$ -Aluminum Cromate;  $\text{ZnCl}_2$ - Zinc Chloride; Nitrous acid- $\text{HNO}_2$ ;  
Hypochlorous acid- $\text{HClO}$ ; hydrochloric- $\text{HCl}$ ;  $\text{H}_2\text{S}$ -hydrosulfuric acid;  $\text{HClO}_4$ -perchloric acid;  $\text{H}_2\text{SO}_4$ -sulfuric acid;  $\text{HNO}_3$ -nitric acid;  $\text{SnO}_2$ -tin dioxide; sulfuric acid- $\text{H}_2\text{SO}_4$ ;

# Chemistry Study Guide Phase Changes Answers

**Iris Stovall,Donald J. DeCoste**

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