



CHEMISTRY

CHEAT SHEET

GAS LAWS



GAS LAWS CHEAT SHEET

Remember: a gas is a substance composed of particles which move from one volume to another (diffusion) and which are separated by spaces. As temperature increases, particles move faster and the pressure increases. As temperature decreases, particles move slower and the pressure decreases.

Boyle's Law
Charles' Law
Avogadro's Law

Boyle's Law	Charles' Law	Avogadro's Law
Pressure (P)	Temperature (T)	Volume (V)
Volume (V)	Pressure (P)	Temperature (T)
Temperature (T)	Volume (V)	Pressure (P)
Pressure (P)	Temperature (T)	Volume (V)

Boyle's Law

Remember: pressure and volume are inversely related.

$$P_1 V_1 = P_2 V_2$$

Remember: as pressure increases, volume decreases. As pressure decreases, volume increases.

Charles' Law

Remember: pressure and temperature are directly related.

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Remember: as temperature increases, volume increases. As temperature decreases, volume decreases.

Avogadro's Law

Remember: volume and number of moles are directly related.

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

Remember: as the number of moles increases, volume increases. As the number of moles decreases, volume decreases.

Ideal Gas Law

The ideal gas law is a combination of Boyle's, Charles', and Avogadro's laws. It is written as follows:

$$PV = nRT$$

Where: P = pressure, V = volume, n = number of moles, R = gas constant, T = temperature.

Remember: as pressure increases, volume decreases. As pressure decreases, volume increases.

Remember: as temperature increases, volume increases. As temperature decreases, volume decreases.

Remember: as the number of moles increases, volume increases. As the number of moles decreases, volume decreases.

$P = \frac{nRT}{V}$	$V = \frac{nRT}{P}$
$n = \frac{PV}{RT}$	$T = \frac{PV}{nR}$

Chemistry Study Guide Gas Laws

Darrell D. Ebbing, Susan M. Schelble



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