

CHEMISTRY



Periodic Table of Elements

- Elemental Forms:**
 - Monatomics:** Noble gases; helium, neon, argon, krypton, xenon, and radon; nonmetals.
 - Diatomics:** Generic fluorine, chlorine, hydrogen, oxygen, and nitrogen; Liquid bromine; Sodium iodide.
 - Molecular solids:** Sulfur
 - Covalent solids:** Carbon (diamond, graphite)
 - Metals:** Most common type of element; high coordination numbers; very dense.
- Types of covalence:** Large variety due to ionic, covalent, and metallic bonding options.
 - Basic:** Electron exchange
 - Covalent:** Shared electrons
 - Polar covalent:** Unequal sharing
 - Ion covalent:** Central ion + Solvent
 - Transition metal covalence:**
 - Coordination covalence:** Non-bonding ligands
 - Organometallics:** Metal-C bond
 - Complexes:** molecules; Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
 - Family of elements:** Columns in periodic table
 - The location of orbitals produces an expanded effect

Periodic Trends

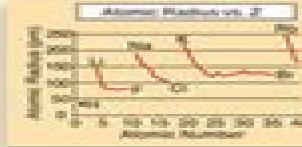
- Chemistry determined by the behavior of valence electrons
- Atoms or ions get larger as you go down the column; electron shells are added
- Cations and anions (same charge) decrease in size as the atomic number (Z) increases

Practical Application

- The periodic table is used to predict the properties of elements in the periodic table.

Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom, $Z = \text{number of electrons}$; for an anion of charge $-n$, $Z = n$ electrons; for a cation of charge $+n$, $Z = n$ electrons
- Mass number (A):** $A = Z + N$, where N = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same Z , different A ; an element can have one or more isotopes; any sample contains a mixture of isotopes (in practical terms, use the average mass number or atomic weight (weighted average of natural isotopes for a given element))



Atomic Quantum Numbers & Orbitals

Pauli principle: An angular momentum (orbital shape) and magnetic (orbital direction) \Rightarrow electron spin

- For each n , the possible l values are $0, 1, \dots, n-1$
- For each l , the possible m_l values are $-l, \dots, l$
- For m_l , three possible values: $+1/2$ and $-1/2$ (spin up, spin down)
- Each value of n electrons is a "shell" in the atomic structure; each l electrons is a subshell
- n -orbital: $l = 0, 1$ type
- l -orbital: $l = 2, 3$ types
- g -orbital: $l = 0, 3$ types
- f -orbital: $l = 3, 7$ types

Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe many-electron atoms and ionic; a more rigorous treatment determines exact energy levels and orbital properties that result in that subshells have different energies; the filling of levels is guided by the Aufbau principle:



| n | l | m_l | ℓ | Subshell |
|-----|-----|-------|--------|----------|
| 1 | 0 | 0 | - | 1s |
| 2 | 0 | 0 | - | 2s |
| 2 | 0 | 0 | - | 2p |
| 2 | 1 | -1 | - | 2p |
| 2 | 1 | 0 | - | 2p |
| 2 | 1 | 1 | - | 2p |
| 3 | 0 | 0 | - | 3s |
| 3 | 1 | -1 | - | 3p |
| 3 | 1 | 0 | - | 3p |
| 3 | 1 | 1 | - | 3p |
| 4 | 0 | 0 | - | 4s |
| 4 | 1 | -1 | - | 4p |
| 4 | 1 | 0 | - | 4p |
| 4 | 1 | 1 | - | 4p |
| 5 | 0 | 0 | - | 5s |
| 5 | 1 | -1 | - | 5p |
| 5 | 1 | 0 | - | 5p |
| 5 | 1 | 1 | - | 5p |

- Electron configuration:** Atomic orbital occupancy for an atom
- Fundamental principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons (one with spin up, $m_l = +1/2$, and one with spin down, $m_l = -1/2$)
- Maxwell's rule:** Electrons fill p , d , and f subshells in a manner maximizing the number of unpaired electrons (half-filled orbitals); spin is random!
- Ionization potential (IP):** Energy required to remove electrons from an atom or ion
 - First IP: Removal of first valence electron
 - Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale; 0 for rare gas

College Chemistry Study Guide

Morris Hein, Susan Arena

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