

Stoichiometry Worksheet # 1

- 1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

- 2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

- 3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

- 4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

- 5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

Ch 12 Stoichiometry Answers

John Moore, Richard H. Langley

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the authors' names.

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Solutions Manual for Chemistry: Molecules Matter and Change, Fourth Edition Julie Henderleiter, Charles Trapp, Peter Atkins, Loretta Jones, 1999-08-13 This student companion is a supplement to Chemistry Molecules Matter and Change 4th edition with CD ROM It features guided reading strategies collaborative learning sheets and strategies for using CD ROM tools

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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

Basic Concepts of Chemistry Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on

outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter WileyPLUS sold separately from text **Programmed Problem-solving for First-year Chemistry** Arnold B. Loebel,1983

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Teaching Science for Understanding James Joseph Gallagher,2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities **Chemical Problem Solving Using Dimensional Analysis** Robert Nakon,1988

Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook McGraw Hill,2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change **McGraw-Hill's SAT Subject Test: Chemistry, 2ed** Thomas Evangelist,2009-02-01 We want to help you score high on the SAT Chemistry test We ve put all of our proven expertise into McGraw Hill s SAT Subject Test Chemistry to make sure you re fully prepared for this difficult exam With this book you ll get essential skill building techniques and strategies created by a leading high school chemistry teacher You ll also get 5 full length practice tests hundreds of sample questions and all the facts about the current exam With McGraw Hill s SAT Subject Test Chemistry we ll guide you step by step through your preparation program and give you the tools you need to succeed 5 full length sample tests with complete explanations for every question 40 top test items to remember on exam day A step by step review of all topics covered on the exam Teacher recommended tips and strategies to help you raise your score *Answer Book to Accompany Boikess/Edelson, Chemical Principles, Third Edition* Robert S. Boikess,1985

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