

Answers to Chapter 9 Study Questions

Study Questions:

1. Why is it important that we eat food and breathe oxygen? Food is broken down through digestion to small organic molecules such as glucose. Glucose and oxygen are reactants in cellular respiration, a process by which cells generates energy (ATP and heat) and produces the waste products (CO_2 and H_2O). Without this energy, cells cease to function and organisms die.
2. Why do plants need CO_2 and H_2O as well as light energy? These molecules are reactants for photosynthesis – the process where plants generate organic molecules (glucose) and oxygen. Plants used these molecules to make ATP via cellular respiration.
3. What types of molecules are used within the body to generate energy? Monosaccharides (glucose), fatty acids, glycerol, and amino acids. These molecules are oxidized to generate CO_2 and water. In general, why does glucose have more potential energy than CO_2 ? Glucose has more chemical bonds that can be rearranged through chemical reactions. The rearrangement of some of these bonds releases energy – thus the bonds in glucose contain much more potential energy. CO_2 does not have chemical bonds that can be easily rearranged and thus it has very little potential energy.
4. Is the oxidation of glucose an endergonic or an exergonic reaction? Exergonic. Why? The change in free energy is -686 kcal/mol. How is the energy that is generated through the oxidation of glucose use in the cell? About 34% of the energy is used to generate ATP. The rest is released as heat. Heat can be use for thermoregulation or dissipated from the body through sweat.
5. Chemical reactions involve the transfer of electrons between molecules. What happens when electrons are relocated between molecules? Energy is released when electrons are transferred. This is why molecules such as methane and glucose have so much potential energy.
6. What are oxidation-reduction reactions? Reactions that involve the partial or complete transfer of one or more electron from one reactant to another. Use the reaction below to answer the following questions.
$$\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O} + \text{energy}$$

What is oxidation and which one of the reactants in the above equation is the reducing agent (electron donor)? Oxidation is the pull of shared electrons away from one atom. In the equation, electrons in methane are equally shared between C and H because these atoms have similar electronegativities. When methane combines with 2 O_2 to produce CO_2 , and H_2O , the electrons are pulled away from C and towards O when CO_2 is formed because O is more electronegative. Thus the carbon in CH_4 , loses shared electrons and it is the reducing agent (electron donor).

What is reduction and which one of the reactants in the above equation is the oxidizing agent (electron acceptor)? Reduction is the pull of shared electrons towards another atom. When methane reacts with 2 O_2 molecules to form CO_2 and H_2O , the electrons that were shared equally in O_2 are now pulled towards O in H_2O . Thus the oxygen in H_2O "gains" the shared electrons and it is the oxidizing agent (electron acceptor).

Why is oxygen such a powerful oxidizing agent? It is very electronegative and will readily accept electrons. What type of redox reaction releases the most chemical energy? Reactions where electrons are relocated to oxygen.

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David G. Myers, Richard O. Straub

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