

Pre-Lab Assignment -- Lab 11 - Chemical Reactions

Answers: Chemical Reactions

1. Which (one or two) of the following chemical reactions are first order reactions?



2. Why do second order reactions (in general) have decreasing reaction rates?

Second-order reactions decrease rapidly as reactants are depleted, so the reaction rate decreases over time. Conversely, second-order reactions increase rapidly as reactants are added, so the reaction rate increases over time.

Chemical equilibrium: Reactions are dynamic processes where reactants and products are in equilibrium.

Reactants and products are in equilibrium when forward and reverse reactions are at the same rate. Equilibrium is represented by the equilibrium constant, K_{eq} , which is the ratio of the concentrations of products to reactants.

Strong equilibrium: Reactions with equilibrium constants equal or near 1000 are considered strong equilibrium reactions. These types of reactions have equilibrium constants of $K_{\text{eq}} > 1000$ or $K_{\text{eq}} \gg 1000$. For strong equilibrium, the equilibrium constant is usually very close to 1000.

Weak equilibrium: Reactions with equilibrium constants less than 1000 are considered weak equilibrium reactions. These types of reactions have equilibrium constants of $K_{\text{eq}} < 1000$. For weak equilibrium, the equilibrium constant is usually very close to 1.

Acid dissociation: Acids dissociate to release protons for basic reactant molecules. These are called conjugate acids.

Conjugate bases are the species that receive the proton. These are called conjugate bases.

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