

## CHEMFILE MINI-GUIDE TO PROBLEM SOLVING

### *Dilutions*

Suppose you work in the laboratory of a paint company where you use 1000 mL of a 0.1 M solution of zinc chloride in a quality-control test that you carry out 10 times a day. It would be tedious and time-consuming to continuously measure out small amounts of  $ZnCl_2$  to make 1000 mL of this solution. Of course, you could make many liters of the solution at one time, but that would require several large containers to store the solution. The answer to this problem is to make a much more concentrated solution and then dilute it with water to make the less concentrated solution that you need. The more-concentrated solution is called a *stock solution*. You could make a 1 M  $ZnCl_2$  solution by measuring out 1 mole of zinc chloride, 136.3 g, and dissolving it in enough water to make a liter of solution. This solution is 10 times as concentrated as the solution you need. Every time you need the test solution, you can measure out 100 mL of the 1 M solution and dilute it to 1000 mL, or yield 1000 mL of 0.1 M  $ZnCl_2$  solution.

To make a solution by dilution, you must determine the volume of stock solution to use and the amount of solvent needed to dilute to the concentration you need. As you have learned, the molarity of a solution is its concentration in moles of solute per liter of solution. Molarity is found by dividing the moles of solute by the number of liters of solution:

$$M = \frac{\text{moles solute}}{\text{liter solution}}$$

So, for a measured volume of any solution:

$$\text{amount of solute in mol} = \text{molarity} \times \text{volume of solution}$$

If this measured volume of solution is diluted to a new volume by adding solvent, the new, larger volume will contain the same number of moles of solute. Therefore, where 1 and 2 represent the concentrated and diluted solutions:

$$\text{molarity}_1 \times \text{volume}_1 = \text{moles solute} = \text{molarity}_2 \times \text{volume}_2$$

Therefore:

$$\text{molarity}_1 \times \text{volume}_1 = \text{molarity}_2 \times \text{volume}_2$$

This relationship applies whenever solution 2 is made from solution 1 by dilution.

# Chemfile Mini Guide

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