

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING**2. PLAN**

- What step is needed to calculate the concentration of the diluted solution?

Apply the principle that $\text{volume}_1 \times \text{molarity}_1 = \text{volume}_2 \times \text{molarity}_2$.

$$M_1V_1 = M_2V_2 \xrightarrow[\text{substitute each of the known quantities for its symbol and calculate}]{\begin{array}{l} \text{rearrange the equation} \\ M_1V_1 = M_2V_2 \\ \text{algebraically to solve for } M_2 \end{array}} M_2 = \frac{M_1V_1}{V_2}$$

$$M_2 = \frac{M_1 \times V_1}{V_2}$$

3. COMPUTE

Note: Even though molarity is moles per liter, you can use volumes in milliliters along with molarity whenever the units cancel.

$$M_2 = \frac{4.74 \text{ M} \times 50.00 \text{ mL}}{250.00 \text{ mL}} = 0.948 \text{ M}$$

4. EVALUATE

- Are the units correct?
- Is the number of significant figures correct?
- Is the answer reasonable?

Yes; molarity (mol/L) was required.
Yes; the number of significant figures is correct because the data had a minimum of three significant figures.
Yes; the computation is the same as 4.74/5, which is a little less than 1.

PRACTICE

1. Complete the table below by calculating the missing value in each row.

Molarity of stock solution	Volume of stock solution	Molarity of dilute solution	Volume of dilute solution	
a. 0.500 M KBr	20.00 mL	7 M KBr	100.00 mL	ans: 0.100 M
b. 1.00 M LiOH	7 mL	0.075 M LiOH	500.00 mL	ans: 38 mL
c. 7 M HI	5.00 mL	0.0493 M HI	100.00 mL	ans: 0.986 M
d. 12.0 M HCl	0.250 L	1.8 M HCl	7 L	ans: 1.7 L
e. 7.44 M NH ₃	7 mL	0.093 M NH ₃	4.00 L	ans: 50. mL

Chemfile Mini Guide To Problem Solving Answers

A. Lamont Tyler, Editors of REA



Chemfile Mini Guide To Problem Solving Answers:

Holt Chemistry File, 1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket

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Problem-Solving Raymond Chang, 2005 By Brandon J Cruickshank Northern Arizona University and Raymond Chang is a success guide written for use with General Chemistry It aims to help students hone their analytical and problem solving skills by presenting detailed approaches to solving chemical problems Solutions for all of the texts even numbered problems

are included **Problem Solving Guide and Workbook for Introductory Chemistry by Steve Russo, Mike Silver**
 Saundra Yancy McGuire, 2007 **Modern Chemistry** Holt Rinehart & Winston, 2005-06-30 **Problem Solving for Chemistry** Edward I. Peters, 1976 **Selected Solutions for Chemical Principles** Kenneth C. Brooks, Steven S. Zumdahl, 1992 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples **Chemistry Problem Solving** James W. Wheeler, 1979-07-01 Problem Solving David W. Brooks, Fred Sicilio, 1972 **Problem Solving in Chemistry** Glen Tilbury, 1976 **Problem Solving in Chemistry** Glen Tilbury, 1984-07-01 **The Answer Key: A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success (Volume Two) (First Edition)** Rachel Turoscy, 2018-08-09 The Answer Key A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success Volume 2 is a concise and accessible textbook that covers the critical information a student needs to understand the basic mathematics used in chemistry courses The book provides easy to understand step by step instructions for solving general chemistry problems The book begins with chapters dedicated to solutions kinetics and liquids solids and phase changes In subsequent chapters the text covers important topics like equilibrium concentrations strong and weak acids and bases the Common Ion Effect and reaction mechanisms It also covers the equilibrium between a solid and its respective ions in a solution as well as the second law of Thermodynamics The text also addresses Gibbs Free Energy equilibrium constants and electrolysis calculations Each chapter contains sample problems and practice problems to help further understanding of how math and chemistry go hand in hand The Answer Key is an excellent resource for any undergraduate course that deals with the basic concepts of general chemistry *Chemical Problem-solving by Dimensional Analysis* Arnold B. Loebel, 1974

Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

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