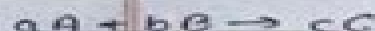


CHEMICAL KINETICS

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Reaction Rate

$$R = \frac{\Delta A}{\Delta t} \quad A_t = -\frac{d[A]}{dt} = \text{slope of conc. - time}$$



$$R_t = \frac{-1}{a} \frac{d[A]}{dt} = \frac{-1}{b} \frac{d[B]}{dt} = \frac{1}{c} \frac{d[C]}{dt}$$

Differential Rate law

Reaction Rate \propto Conc. of React.

$$\text{Rate} \propto [A]^m [B]^n$$

$$R = k[A]^m [B]^n$$

$$\text{or } \frac{1}{a} \frac{d[A]}{dt} = k[A]^m [B]^n$$

Rate Constant

$[A] = [B] = 1M$ then Rate = k .

Order

$(m+n) = \text{overall order}$

• Zero Order Rxn

Rate independent of conc. of react^{ts}

$$k = \frac{[A]_0 - [A]}{t} \quad \text{time}^{-1}$$

$$k = -\text{slope}$$



$$t_{1/2} = \frac{[A]_0}{2k}$$

• First Order Rxn

Rate depends linearly with conc.

$$k = \frac{1}{t} \ln \frac{[A]_0}{[A]} = \frac{2.303}{t} \log \frac{a}{a-x}$$

$$t_{1/2} = \frac{0.693}{k}$$

$$\ln[A] = -kt + \ln[A]_0$$

$$\text{or } k = -\text{slope} = \frac{1}{t} \ln \frac{[A]_0}{[A]}$$

$$\log[A] = \log[A]_0 - \frac{kt}{2.303}$$

$$k = \frac{-\text{slope}}{2.303} = \frac{2.303}{t} \log \frac{[A]_0}{[A]}$$

• For plot b/w $\log \frac{[A]_0}{[A]}$ v/s time

$$\text{slope} = \frac{k}{2.303}$$

• (gas Phase 2nd order Rxn

$$k = \frac{2.303 \log \frac{P_i - P_f}{P_f - P_i}}{t}$$

$$= \frac{2.303 \log \frac{P_i}{2P_i - P_f}}{t}$$

• Temp. Dependence

$$E_a = E(\text{threshold}) - E(\text{reactant molecule})$$

• Arrhenius Eqⁿ

$$k = A e^{-E_a/RT}$$

- Exothermic: $\Delta E = -ve$

- Endothermic: $\Delta E = +ve$

$$\log k = \log A - \frac{E_a}{2.303R} \times \frac{1}{T}$$

$$\text{slope} = \frac{-E_a}{2.303R}$$

$$\log \frac{k_1}{k_2} = \frac{E_a}{2.303R} \left[\frac{1}{T_2} - \frac{1}{T_1} \right]$$

Chemical Kinetic Solution Manual

Raymond Chang



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web documents to be submitted 1 the personal information provided in this form will be used solely for the purpose of transfer request 2 the school will not disclose the reasons for selection or non selection of student 3 this application is deemed to be unsuccessful if you do not hear from the school

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