

## Experiment 17-Chemical Reactions Lab

Since the Middle Ages, when ancient physicians attempted to find a magical substance that would cure all diseases, humans have been fascinated with chemical reactions. In order to effectively describe these changes, a system for and series of short hand notation was developed to aid ancient and now modern chemists classify chemical reactions.

The names of the basic building blocks of matter (elements) have historically interesting names such as Ytterbium, named after a cute little town in Sweden, and Plumbus, named after a metal (lead) that was used for water pipes in ancient Rome. The names of *compounds*, or combinations of different elements, are much more complicated. To simplify matters, chemists use one or two letter abbreviations for the elements. Writing a compound is simplified by using a *chemical formula*, or a series of symbols that represent the elements present in the compound.

Chemical compounds, as we will learn, have definite, whole number ratios of various. Typically, the atoms in a compound are listed in the formula in the order (from left to right) that the elements appear on the periodic table. The number ratios of the atoms in a compound are shown by using *subscripts* (little numbers that are lower than the rest of the text) just to the right of the element. Examples of formulas are shown below.

NaCl	sodium chloride, with 1 sodium atom and 1 chlorine atom
CO <sub>2</sub>	carbon dioxide, with 1 carbon atom for every 2 oxygen atoms
HCl	hydrochloric acid, with 1 hydrogen atom for each chlorine atom

A chemical reaction involves elements and/or compounds combining, separating or changing partners in various ways. Again, describing such a complicated process in words is difficult, so a *chemical reaction equation* is used. The equation shows the chemical formula(s) of the starting compound(s), called *reactants*, and the final compound(s), called *products*. The reactants are always shown on the left, and the products on the right, and they are separated by an arrow to ensure that no one is confused on which direction the reaction is going. A general chemical reaction equation for a reaction called combination is shown below, and a specific example follows.



If you would like to read the equation aloud, this is what the general reaction #1 would say: A reacts with B to make C. Equation 2 would read: Carbon reacts with oxygen to make carbon dioxide. The funny letters in parentheses in equation 2 represent the *physical state* that the matter is in, such as gas (g), liquid (l), solid (s), or dissolved in water (aq).

In order to aid chemists in recognizing and describing reactions, several types of reactions are classified into groups. Equation 2 is an example of a *combination* reaction, where two reactants combine to form 1 product. Examples of other types of reactions are listed below:

*Decomposition*, where one reactant decomposes or separates into 2 or more simpler products:



Carbonic acid decomposes in water to make carbon dioxide gas and water.

# Chemical Reaction Lab Report Conclusion

**Dorin, Prentice Hall Dictionary Editors**



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