

Chemistry 110 Study Guide Chapters 55-57

Chapter 55

Key Concepts:

- 1. The electronic structure of atoms determines their chemical properties.*
- 2. The types of elements are determined by the number of protons in the nucleus.*
- 3. The number of protons in an element's nucleus determines how strong the atom's nucleus is held together (termed "atomic number").*
- 4. The types of elements are determined by the number of electrons in the outermost energy level of the atom (termed "atomic number").*
- 5. Electrons in atoms are arranged in shells around the nucleus.*
- 6. The shells are filled in order of increasing energy.*
- 7. The first shell can hold a maximum of 2 electrons, the second shell can hold a maximum of 8 electrons, and so on.*
- 8. The outermost shell is called the valence shell.*
- 9. The valence shell determines how an atom will interact with other atoms.*
- 10. Atoms interact with each other in order to achieve a full valence shell.*
- 11. The types of elements are determined by the number of valence electrons in the outermost shell of the atom.*
- 12. The number of valence electrons determines how an atom will interact with other atoms.*
- 13. Atoms with the same number of valence electrons will interact with each other in the same way.*

Important Events:

- 1. The discovery of the electron in 1897 by J. J. Thomson.*
- 2. The discovery of the nucleus in 1911 by Ernest Rutherford.*
- 3. The discovery of the proton in 1918 by Ernest Rutherford.*
- 4. The discovery of the neutron in 1932 by James Chadwick.*
- 5. The development of the periodic table of the elements by Dmitri Mendeleev in 1869.*

Notable Figures:

- 1. J. J. Thomson - discovered the electron in 1897.*
- 2. Ernest Rutherford - discovered the nucleus in 1911 and the proton in 1918.*

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

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