

Chemistry 110 Study Guide Chapters 55-57

Chapter 55

Key Concepts:

1. The electronic structure of atoms determines their chemical properties.
2. The types of elements are determined by the number of protons in the nucleus.
3. The number of protons in an element's nucleus determines how strong the atom's nucleus is held together (termed "atomic number").
4. The types of elements are determined by the number of electrons in the outermost energy level of the atom (termed "atomic number").
5. Electrons in atoms are arranged in shells around the nucleus.
6. The shells are filled in order of increasing energy.
7. The first shell can hold a maximum of 2 electrons, the second shell can hold a maximum of 8 electrons, and so on.
8. The outermost shell is called the valence shell.
9. The valence shell determines how an atom will interact with other atoms.
10. Atoms interact with each other in order to achieve a full valence shell.
11. The types of elements are determined by the number of valence electrons in the outermost shell of the atom.
12. The number of valence electrons determines how an atom will interact with other atoms.
13. Atoms with the same number of valence electrons will interact with each other in the same way.

Important Events:

1. The discovery of the electron in 1897 by J. J. Thomson.
2. The discovery of the nucleus in 1911 by Ernest Rutherford.
3. The discovery of the proton in 1918 by Ernest Rutherford.
4. The discovery of the neutron in 1932 by James Chadwick.
5. The development of the periodic table of the elements by Dmitri Mendeleev in 1869.

Notable Figures:

1. J. J. Thomson - discovered the electron in 1897.
2. Ernest Rutherford - discovered the nucleus in 1911 and the proton in 1918.

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