

Chemistry diagnostic test (Topic 12)

1. B – the p^2 config. of S is less stable, electron- electron repulsion.



2. A – B is lower than Be because two s electrons shield the 2p electrons

3. E – Au and Ag both have full d orbitals(10 electrons) (s^1)

4. E – electronegativity increases from lower left to upper right (F-most electronegative)

5. B – sulfur's oxidation numbers: 2-, 2+, 4+, 6+

6. A – atomic radii increase right to left and top to bottom

7. C – cations are smaller, the higher the charge the smaller the cation

8. A – Group 1 has low 1st ionization energy, but 2nd is pulling an inner core electron(big jump)

9. E – equal moles - 0.25

10. D – greatest molar mass

11. D – greatest mass/volume

12. E – all same temp

13. E – Phosphorus – high reactivity with oxygen not found free in nature

14. B – not all have a 2+ oxidation

15. A – all substances contain London dispersion forces

16. D – hydrogen must be bonded to nitrogen, oxygen or fluorine

17. A – propane (molar mass – 44) dispersion forces only – lowest temp
Ethanol (molar mass – 46) also has hydrogen bonding
Water (molar mass – 18) very strong hydrogen bonds

18. C

19. A – balanced equation is
 $Na_2CO_3 + 2HCl \rightarrow 2NaCl + CO_2 + H_2O$
Limiting is sodium carbonate

20. C – $(0.300 \text{ moles/L})(.300\text{L}) = 0.0900 \text{ mol Cl}^- \text{ ion}$ – requires 1/3 as much aluminum chloride or 0.0300 mol
 $0.0300 \text{ mol}/0.500\text{M} \text{ mol/mol/L} = 0.0600\text{L} = 60.0 \text{ ml}$

21. B - $0.500 \text{ M} \times 0.250\text{L} \times 200\text{g/mol} = 25.0 \text{ g}$

22. D- $2 C_3H_8O + 9O_2 \rightarrow 6CO_2 + 8H_2O$

23. C – 3.00 carbon = $\frac{1}{3}$ mole of C therefore $\frac{1}{3}$ mol of Z combines
 $0.5 \text{ mol Z} \times 36.0 \text{ g/mol}(3\text{timesC}) = 18.0 \text{ g Z}$

24. E – 4 moles of O

25. C – an alkali metal reacts with water to produce hydrogen gas and an aqueous alkali metal hydroxide

26. D – $H_2SO_4 + NaOH \rightarrow NaHSO_4 + H_2O$ (diprotonic acid and monobasic base)

27. E – X has a 3+ charge

28. E

29. B – 1 mole of magnesium requires 2 moles of HCl therefore HCl is limiting

30. D – a formation reaction produces 1 mole of a product

Chemistry Answer Key Diagnostic Test Topic 4

Adam Robinson, John Katzman

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Table of Contents Chemistry Answer Key Diagnostic Test Topic 4

1. Understanding the eBook Chemistry Answer Key Diagnostic Test Topic 4
 - The Rise of Digital Reading Chemistry Answer Key Diagnostic Test Topic 4
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Answer Key Diagnostic Test Topic 4
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Answer Key Diagnostic Test Topic 4
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Answer Key Diagnostic Test Topic 4
 - Personalized Recommendations
 - Chemistry Answer Key Diagnostic Test Topic 4 User Reviews and Ratings
 - Chemistry Answer Key Diagnostic Test Topic 4 and Bestseller Lists
5. Accessing Chemistry Answer Key Diagnostic Test Topic 4 Free and Paid eBooks
 - Chemistry Answer Key Diagnostic Test Topic 4 Public Domain eBooks
 - Chemistry Answer Key Diagnostic Test Topic 4 eBook Subscription Services
 - Chemistry Answer Key Diagnostic Test Topic 4 Budget-Friendly Options
6. Navigating Chemistry Answer Key Diagnostic Test Topic 4 eBook Formats

- ePub, PDF, MOBI, and More
- Chemistry Answer Key Diagnostic Test Topic 4 Compatibility with Devices
- Chemistry Answer Key Diagnostic Test Topic 4 Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Chemistry Answer Key Diagnostic Test Topic 4
- Highlighting and Note-Taking Chemistry Answer Key Diagnostic Test Topic 4
- Interactive Elements Chemistry Answer Key Diagnostic Test Topic 4

8. Staying Engaged with Chemistry Answer Key Diagnostic Test Topic 4

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Answer Key Diagnostic Test Topic 4

9. Balancing eBooks and Physical Books Chemistry Answer Key Diagnostic Test Topic 4

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chemistry Answer Key Diagnostic Test Topic 4

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chemistry Answer Key Diagnostic Test Topic 4

- Setting Reading Goals Chemistry Answer Key Diagnostic Test Topic 4
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Answer Key Diagnostic Test Topic 4

- Fact-Checking eBook Content of Chemistry Answer Key Diagnostic Test Topic 4
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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