

Unit Three: Study Guide --- Chemistry

Answer the following questions to prepare for your next test. The completed study guide is due Test Day (Friday October 30). Be sure to get extra help if you struggle on any of the concepts in this study guide. All answers should be on a separate sheet of paper.

1. Define the following words.

- a. atomic size
- b. ionization energy
- c. ions
- d. cations
- e. anions
- f. ionic compounds
- g. covalent compounds
- h. electron configurations
- i. valence electrons
- j. energy levels

2. Write the FULL electron configurations for the following elements.

- a. Sodium
- b. Potassium
- c. Rubidium
- d. Neon
- e. Helium
- f. Cobalt
- g. Gallium
- h. Argon
- i. Chlorine
- j. Carbon
- k. Boron
- l. Nitrogen

3. Identify the elements represented by the following electron configurations.

- a. $1s^2 2s^1$
- b. $1s^2 2s^2 2p^1 3s^1$
- c. $1s^2 2s^2 2p^2$
- d. $1s^2 2s^2 2p^3 3s^1 3p^1$
- e. $1s^2 2s^2 2p^3 3s^1 3p^1 4s^1 3d^1$
- f. $1s^2 2s^2 2p^3 3s^1 3p^1 4s^1 3d^1 4p^1$
- g. $1s^2 2s^2 2p^3 3s^1 3p^1 4s^1 3d^1 4p^1 5s^1$
- h. $[Ar] 4s^2$
- i. $[He] 2s^2 2p^1$
- j. $[Ne] 2s^2 2p^6 3s^1 3p^6 4s^1 3d^1 4p^1$
- k. $[Ar] 3s^2$
- l. $[Ar] 3s^2 3p^1 3s^1 3p^1 4s^1 3d^1 4p^1$
- m. $[Ar] 3s^1$
- n. $[Ar] 3s^2 3p^1 3s^1 3p^1 4s^1 3d^1$

4. Identify whether the trend is increasing, decreasing, or stays the same.

- a. ionization energy as you move left to right across the table
- b. ionization energy as you move down the table
- c. ionization energy as you move down the table
- d. atomic radius as you move left to right on the table
- e. atomic radius as you move down the table
- f. ionization energy as you move left to right across the table
- g. atomic size as you move down the periodic table
- h. atomic size as you move across (left to right) the periodic table

5. Complete the table below.

Compound Formula	Ionic or Covalent?	Compound Name
NaCl		

Chemistry Chemical Bonding Study Guide Answers

Susan McMurry

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Table of Contents Chemistry Chemical Bonding Study Guide Answers

1. Understanding the eBook Chemistry Chemical Bonding Study Guide Answers
 - The Rise of Digital Reading Chemistry Chemical Bonding Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chemical Bonding Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chemical Bonding Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chemical Bonding Study Guide Answers
 - Personalized Recommendations
 - Chemistry Chemical Bonding Study Guide Answers User Reviews and Ratings
 - Chemistry Chemical Bonding Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Chemical Bonding Study Guide Answers Free and Paid eBooks

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- Adjustable Fonts and Text Sizes of Chemistry Chemical Bonding Study Guide Answers
- Highlighting and Note-Taking Chemistry Chemical Bonding Study Guide Answers
- Interactive Elements Chemistry Chemical Bonding Study Guide Answers

8. Staying Engaged with Chemistry Chemical Bonding Study Guide Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Chemical Bonding Study Guide Answers

9. Balancing eBooks and Physical Books Chemistry Chemical Bonding Study Guide Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chemistry Chemical Bonding Study Guide Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chemistry Chemical Bonding Study Guide Answers

- Setting Reading Goals Chemistry Chemical Bonding Study Guide Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Chemical Bonding Study Guide Answers

- Fact-Checking eBook Content of Chemistry Chemical Bonding Study Guide Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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