

Unit Three Study Guide -- Chemistry

Answer the following questions to prepare for your next test. The completed study guide is due Test Day (Friday October 3). Be sure to get extra help if you struggle on any of the concepts in this study guide. All answers should be on a separate sheet of paper.

1. Define the following words.

- a. atomic size
- b. ionization energy
- c. ions
- d. cations
- e. anions

- f. ionic compounds
- g. covalent compounds
- h. electron configuration
- i. valence electrons
- j. energy levels

2. Write the FULL electron configuration for the following elements.

- a. Sodium
- b. Potassium
- c. Rubidium
- d. Vanadium
- e. Helium
- f. Cobalt

- g. Gallium
- h. Argon
- i. Chlorine
- j. Carbon
- k. Bromine
- l. Krypton

3. Identify the element represented by the following electron configurations.

- a. $1s^2 2s^1$
- b. $1s^2 2s^2 2p^4 3s^1$
- c. $1s^2 2s^2 2p^2$
- d. $1s^2 2s^2 2p^6 3s^2 3p^1$
- e. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^1$
- f. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1$
- g. $1s^2 2s^2 2p^2$

- h. $[Ar] 4s^2 3d^{10}$
- i. $[He] 2s^2 2p^2$
- j. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^{10} 4p^2$
- k. $[Xe] 6s^2 5d^1$
- l. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^{10} 4p^2$
- m. $[Kr] 5s^1$
- n. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^1$

4. Identify whether the trend is increasing, decreasing, or stays the same.

- a. electronegativity as you move left to right across the table
- b. ionization energy as you move down the table
- c. electronegativity as you move down the table
- d. atomic radius as you move left to right on the table
- e. atomic radius as you move down the table
- f. ionization energy as you move left to right across the table
- g. atomic size as you move down the periodic table
- h. atomic size as you move across (left to right) the periodic table

5. Complete the table below.

Compound Formula	Ionic or Covalent?	Compound Name
H_2O		

Chemistry Chemical Bonding Study Guide Answers

Susan McMurry



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