

## Chemistry

## FORMULAE SHEET

$$n = \frac{m}{MM}$$

$$q = mc\Delta T$$

$$pK_a = -\log_{10}[K_a]$$

$$c = \frac{n}{V}$$

$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$$

$$A = \epsilon/c = \log_{10} \frac{I_0}{I}$$

$$PV = nRT$$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

Avogadro constant,  $N_A$  .....  $6.022 \times 10^{23} \text{ mol}^{-1}$

Volume of 1 mole ideal gas: at 100 kPa and  
at 0°C (273.15 K) ..... **22.71 L**  
at 25°C (298.15 K) ..... **24.79 L**

Gas constant ..... **8.314 J mol<sup>-1</sup> K<sup>-1</sup>**

Ionisation constant for water at 25°C (298.15 K),  $K_w$  .....  $1.0 \times 10^{-14}$

Specific heat capacity of water ..... **4.18 × 10<sup>3</sup> J kg<sup>-1</sup> K<sup>-1</sup>**

## DATA SHEET

## Solubility constants at 25°C

<i>Compound</i>	$K_{sp}$ 10 <sup>-9</sup>	<i>Compound</i>	$K_{sp}$ 10 <sup>-6</sup>
Barium carbonate	2.58 × 10 <sup>-9</sup>	Lead(II) bromide	6.60 × 10 <sup>-6</sup>
Barium hydroxide	2.55 × 10 <sup>-4</sup>	Lead(II) chloride	1.70 × 10 <sup>-5</sup>
Barium phosphate	1.3 × 10 <sup>-25</sup>	Lead(II) iodide	9.8 × 10 <sup>-9</sup>
Barium sulfate	1.08 × 10 <sup>-10</sup>	Lead(II) carbonate	7.40 × 10 <sup>-14</sup>
Calcium carbonate	3.36 × 10 <sup>-9</sup>	Lead(II) hydroxide	1.43 × 10 <sup>-15</sup>
Calcium hydroxide	5.02 × 10 <sup>-6</sup>	Lead(II) phosphate	8.0 × 10 <sup>-13</sup>
Calcium phosphate	2.07 × 10 <sup>-29</sup>	Lead(II) sulfate	2.53 × 10 <sup>-8</sup>
Calcium sulfate	4.93 × 10 <sup>-5</sup>	Magnesium carbonate	6.82 × 10 <sup>-6</sup>
Copper(II) carbonate	1.4 × 10 <sup>-10</sup>	Magnesium hydroxide	5.61 × 10 <sup>-12</sup>
Copper(II) hydroxide	2.2 × 10 <sup>-20</sup>	Magnesium phosphate	1.04 × 10 <sup>-24</sup>
Copper(II) phosphate	1.40 × 10 <sup>-37</sup>	Silver bromide	5.35 × 10 <sup>-13</sup>
Iron(II) carbonate	3.13 × 10 <sup>-11</sup>	Silver chloride	1.77 × 10 <sup>-10</sup>
Iron(II) hydroxide	4.87 × 10 <sup>-17</sup>	Silver carbonate	8.46 × 10 <sup>-12</sup>
Iron(III) hydroxide	2.79 × 10 <sup>-39</sup>	Silver hydroxide	2.0 × 10 <sup>-8</sup>
Iron(III) phosphate	9.91 × 10 <sup>-16</sup>	Silver iodide	8.52 × 10 <sup>-17</sup>
		Silver phosphate	8.89 × 10 <sup>-17</sup>
		Silver sulfate	1.20 × 10 <sup>-5</sup>

# Chemistry Data Sheet For Ch1fp

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  4. In chapter 3, this book will examine the practical applications of Chemistry Data Sheet For Ch1fp in daily life. The third chapter will showcase real-world examples of how Chemistry Data Sheet For Ch1fp can be effectively utilized in everyday scenarios.
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