

# Equilibria Test SCH4U Name:

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Multiple Choice: Circle the most correct answer. (15%)

1. Which of the following expressions is the correct equilibrium-constant expression for the reaction below?



- A)  $[\text{HF}][\text{H}_2\text{O}] / [\text{H}_3\text{O}^+][\text{F}^-]$  B)  $1 / [\text{HF}]$  C)  $[\text{H}_3\text{O}^+][\text{F}^-] / [\text{HF}][\text{H}_2\text{O}]$   
 D)  $[\text{H}_3\text{O}^+][\text{F}^-] / [\text{HF}]$  E)  $[\text{F}^-] / [\text{HF}]$

2. The  $K_{eq}$  for the equilibrium between  $750 \times 10^3$  at 480.0 °C.



What is the value of  $K_{eq}$  at this temperature for the following reaction?



- A) 0.0752 B)  $5.88 \times 10^{-4}$  C) 0.274 D) 0.0076 E) 0.250

3. At 500.0 K, the equilibrium constant ( $K_p$ ) for the following reaction is 0.345.



atm. The partial pressure of  $\text{SO}_3$  is \_\_\_\_\_ atm.

- A) 82.0 B)  $4.31 \times 10^{-4}$  C) 190 D)  $4.39 \times 10^{-4}$  E) 40.2

4. Phosphorus trichloride and phosphorus pentachloride equilibrate in the presence of molecular chlorine according to the reaction:  $\text{PCl}_3(g) + \text{Cl}_2(g) \rightleftharpoons \text{PCl}_5(g)$

An equilibrium mixture at 400 K contains:  $\text{PPCl}_3 = 0.202$  atm,  $\text{PCl}_3 = 0.256$  atm, and  $\text{PPCl}_5 = 3.45$  atm.

What is the value of  $K_p$  at this temperature?

- A) 66.7 B)  $1.50 \times 10^{-4}$  C)  $1.78 \times 10^{-4}$  D) 2.69 E) 7.54

5. The reaction,  $\text{PCl}_3(g) + \text{Cl}_2(g) \rightleftharpoons \text{PCl}_5(g)$   $K=5.41$

has the  $K_{eq}$  above. If a reaction mixture has the concentrations as shown, in what direction is the reaction proceeding?

Concentrations:  $[\text{PCl}_3] = 0.16 \times 10^{-3} \text{ M}$   $[\text{Cl}_2] = 1.16 \times 10^{-3} \text{ M}$   $[\text{PCl}_5] = 1.12 \times 10^{-3} \text{ M}$

- A) Right to left B) Left to Right C) It is at equilibrium

6. For a solution in a closed container at equilibrium, which of the following pairs represent processes that must be occurring at equal rates?

- A) evaporation and crystallization B) crystallization and sublimation  
 C) fusion and boiling D) evaporation and sublimation  
 E) evaporation and condensation

# Chemistry Equilibrium Test Answers

**Thomas Evangelist**



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*Partial Solutions Manual* Darrell D. Ebbing,David Bookin,1998

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