

**Solubility Rules:**- $\text{NO}_3^-$  all soluble- $\text{Cl}^-$  all soluble except  $\text{AgCl}$ ,  $\text{Hg}_2\text{Cl}_2$  and  $\text{PbCl}_2$ - $\text{SO}_4^{2-}$  all soluble except  $\text{BaSO}_4$ ,  $\text{PbSO}_4$  and  $\text{SrSO}_4$ - $\text{OH}^-$  all insoluble except Group 1,  $\text{Ba}(\text{OH})_2$ ,  $\text{Sr}(\text{OH})_2$ ,  $\text{Ca}(\text{OH})_2$ - $\text{CO}_3^{2-}$  all insoluble except  $\text{NH}_4^+$  and Group 1- $\text{S}^{2-}$  all insoluble except Group 1 and Group 2 and  $\text{NH}_4^+$ **Polyatomics:**-Acetate ( $\text{C}_2\text{H}_3\text{O}_2^-$ )-Chlorate ( $\text{ClO}_3^-$ )-Chlorite ( $\text{ClO}_2^-$ )-Cyanide ( $\text{CN}^-$ )-Hydroxide ( $\text{OH}^-$ )-Nitrate ( $\text{NO}_3^-$ )-Nitrite ( $\text{NO}_2^-$ )-Nitrite ( $\text{NO}_2^-$ )-Carbonate ( $\text{CO}_3^{2-}$ )-Chromate ( $\text{CrO}_4^{2-}$ )-Dichromate ( $\text{Cr}_2\text{O}_7^{2-}$ )-Oxalate ( $\text{C}_2\text{O}_4^{2-}$ )-Sulfate ( $\text{SO}_4^{2-}$ )-Sulfite ( $\text{SO}_3^{2-}$ )-Arsenate ( $\text{AsO}_4^{3-}$ )-Arsenite ( $\text{AsO}_3^{3-}$ )-Phosphate ( $\text{PO}_4^{3-}$ )-Phosphite ( $\text{PO}_3^{3-}$ )-Ammonium ( $\text{NH}_4^+$ )-Bicarbonate ( $\text{HCO}_3^-$ )**Acids:**- $\text{HCl}$  = hydrochloric- $\text{H}_2\text{SO}_4$  = sulfuric- $\text{HNO}_3$  = nitric- $\text{CH}_3\text{COOH}$  = acetic- $\text{H}_3\text{PO}_4$  = phosphoric- $\text{H}_2\text{CO}_3$  = carbonicMatter is the physical material of the universe and exists in 3 states: solid, liquid, gasSubstance is matter that has a uniform and definite compositionMixture is a physical blend of substancesHomogenous Mixture has uniform composition throughout (solution) Heterogeneous Mixture does notPhysical Properties: observed or measuredChemical Properties: ability to undergo chemical changeLaw of Conservation of Mass: matter/mass can be neither created nor destroyedExtensive Properties: mass and volumeIntensive Properties: compositionSolid: definite shape, volume, not easily compressed Liquid: indefinite shape, definite volume, not easily compressedGas: indefinite shape, volume, easily compressed \*Particles are always moving in S, L, GIndications of Chemical Reaction: dissolving, releasing/absorption of heat, production of gas, color change, precipitateQualitative give descriptions, no numbersQuantitative number-basedAccuracy is how close to a single measurementPrecision is how close several measurements are togetherError =  $|\text{expt} - \text{accepted}|$  % Error =  $|\text{expt} - \text{accepted}| / \text{accepted} \cdot 100$  Percent Yield =  $\text{expt} / \text{theoretical} \cdot 100$ 

1 kilo(unit) = 1000 (unit) 1 (unit) = 1000 milli

1 (unit) = 100 centi

micro =  $10^6$ nano =  $10^9$  $K = C + 273$  $C = K - 273$  $\text{H}_2\text{O}$  FP = 273K

BP = 373K Density = mass/volume

Density of  $\text{H}_2\text{O}$  = 1 g/mLSig Figs: For addition & subtraction use least # of decimal places

/ For Multiplication &amp; Division use least # of sig figs

Dalton's Atomic Theory: 1) all elements are composed of atoms

2) atoms of same element are identical

3) atoms can physically mix/chemically combine in whole #d ratios

4) chemical rxns happen; atoms don't change

Proton = pos., change = radioactivityNeutron = neutral, change = isotopeElectron = neg., change = ion

# Chemistry Final Content Review Packet

**Sebastian Brünink**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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