

Solubility Rules:**-NO₃⁻²** all soluble**-Cl⁻²** all soluble except AgCl, Hg₂Cl₂ and PbCl₂**-SO₄⁻²** all soluble except BaSO₄, PbSO₄ and SrSO₄**-OH⁻²** all insoluble except Group 1, Ba(OH)₂, Sr(OH)₂, Ca(OH)₂**-CO₃⁻²** all insoluble except NH₄⁺ and Group 1**-S⁻²** all insoluble except Group 1 and Group 2 and NH₄⁺**Polyatomics:****-Acetate (C₂H₃O₂⁻²)****-Carbonate (CO₃⁻²)****-Arsenate (AsO₄⁻²)****Acids:****-HCl = hydrochloric****-Chlorate (ClO₃⁻¹)****-Chromate (CrO₄⁻²)****-Arsenite (AsO₃⁻²)****-H₂SO₄ = sulfuric****-Chlorite (ClO₄⁻¹)****-Dichromate (Cr₂O₇⁻²)****-Phosphate (PO₄⁻³)****-HNO₃ = nitric****-Cyanide (CN⁻¹)****-Oxalate (C₂O₄⁻²)****-Phosphite (PO₃⁻³)****-CH₃COOH = acetic****-Hydroxide (OH⁻¹)****-Sulfate (SO₄⁻²)****-Ammonium (NH₄⁺)****-H₃PO₄ = phosphoric****-Nitrate (NO₃⁻²)****-Sulfite (SO₃⁻²)****-Bicarbonate (HCO₃⁻²)****-H₂CO₃ = carbonic****-Nitrite (NO₂⁻¹)****Matter** is the physical material of the universe and exists in 3 states: solid, liquid, gas**Substance** is matter that has a uniform and definite composition**Mixture** is a physical blend of substances**Homogeneous Mixture** has uniform composition throughout (solution) **Heterogeneous Mixture** does not**Physical Properties:** observed or measured**Chemical Properties:** ability to undergo chemical change**Law of Conservation of Mass:** matter/mass can be neither created nor destroyed**Extensive Properties:** mass and volume**Intensive Properties:** composition**Solid:** definite shape, volume, not easily compressed **Liquid:** indefinite shape, definite volume, not easily compressed**Gas:** indefinite shape, volume, easily compressed

*Particles are always moving in S, L, G

Indications of Chemical Reaction: dissolving, releasing/absorption of heat, production of gas, color change, precipitate**Qualitative:** give descriptions, no numbers**Quantitative:** number-based**Accuracy:** is how close to a single measurement**Precision:** is how close several measurements are together

Error = |expt - accepted| % Error = |expt-accepted|/accepted * 100 Percent Yield = expt/theoretical * 100

1 kilo(unit) = 1000 (unit) 1 (unit) = 1000 milli

1 (unit) = 100 centi

micro=10⁻⁶nano=10⁻⁹

K = C + 273

C = K - 273

H₂O FP=273K

BP=373K Density = mass/volume

Density of H₂O = 1 g/mL**Sig Figs:** For addition & subtraction use least # of decimal places

/ For Multiplication & Division use least # of sig figs

Dalton's Atomic Theory: 1) all elements are composed of atoms

2) atoms of same element are identical

3) atoms can physically mix/chemically combine

in whole #d ratios

4) chemical rxns happen; atoms don't change

Proton = pos., charge = radioactivity**Neutron** = neutral, charge = isotope**Electron** = neg., charge = ion

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Sebastian Brünink

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