

CHEMISTRY FINAL EXAM REVIEW GUIDE 2012 Scott

DISCLAIMER: THIS IS A SAMPLE OF QUESTIONS/PROBLEMS FROM THIS YEAR. PLEASE BE SURE TO REVIEW YOUR NOTES, LABS, COJ'S, OBJECTIVE SHEETS, ETC. DO NOT RELY ONLY ON THIS REVIEW GUIDE.

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- 1) Which of the following observations is quantitative?
 - a. The liquid turns blue litmus paper red.
 - b. The liquid boils at 100°C.
 - c. The liquid tastes bitter.
 - d. The liquid is cloudy.

- 2) The symbols for units of length in order from largest to smallest are
 - a. m, cm, mm, km.
 - b. mm, m, cm, km.
 - c. km, mm, cm, m.
 - d. km, m, cm, mm.

- 3) A volume of 1 milliliter is equivalent to
 - a. 1 cubic centimeter.
 - b. 1 gram.
 - c. 1 liter.
 - d. 10^{-3} cubic decimeters.

- 4) Which pair of quantities determines the density of a material?
 - a. mass and weight
 - b. volume and weight
 - c. volume and concentration
 - d. volume and mass

- 5) A sample of gold has a mass of 96.5 g and a volume of 5.00 cm³. The density of gold is
 - a. 0.0518 g/cm³.
 - b. 19.3 g/cm³.
 - c. 101.5 g/cm³.
 - d. 48.3 g/cm³.

- 6) The number of grams equal to 0.5 kg is
 - a. 0.0005.
 - b. 0.005.
 - c. 500.
 - d. 5000.

- 7) If 1 inch equals 2.54 cm, how many centimeters equal 1 yard?
 - a. 0.0706 cm
 - b. 14.2 cm
 - c. 30.5 cm
 - d. 91.4 cm

- 8) A chemist obtained the values 5.2246 g, 5.2353 g, and 5.2501 g for the mass of a sample. Without knowing the true mass of the sample, it can be said that these values have
 - a. good precision.
 - b. good accuracy.
 - c. poor precision.
 - d. poor accuracy.

- 9) The measurement 0.035550 g rounded off to two significant figures would be
 - a. 0.03 g.
 - b. 0.35 g.
 - c. 0.036 g.
 - d. 3.5×10^2 g.

- 10) When 64.4 is divided by 2.00, the correct number of significant figures in the result is

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