

## Intro to Chemistry Final Exam Study Guide

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. The Fahrenheit temperature scale is defined by which of the following temperatures?
- Water boils at 100 degrees and freezes at 32 degrees.
  - Water boils at 212 degrees and freezes at 32 degrees.
  - Water evaporates at 212 degrees and freezes at 0 degrees.
  - Liquid water turns to a gas at 100 degrees and to a solid at 0 degrees.
- \_\_\_\_\_ 2. A cold-blooded reptile basks on a warm rock in the sun. Its body is warmed by
- radiation.
  - conduction.
  - convection.
  - Both (a) and (b).
- \_\_\_\_\_ 3. Temperature is a measure of the average \_\_\_\_\_ energy of an object's particles.
- mechanical
  - kinetic
  - potential
  - light
- \_\_\_\_\_ 4. The temperature at which all molecular motion stops is
- 0°C.
  - 0°F.
  - 0 K.
  - All of the above
- \_\_\_\_\_ 5. Convert 500°F to degrees Celsius.
- 260°C
  - 296°C
  - 842°C
  - 958°C
- \_\_\_\_\_ 6. Convert 300 K to the Celsius scale.
- 0°C
  - 27°C
  - 73°C
  - 573°C
- \_\_\_\_\_ 7. Convert 458 K to the equivalent Fahrenheit temperature.
- 185°F
  - 237°F
  - 365°F
  - 731°F
- \_\_\_\_\_ 8. Which of the following temperatures is impossible to measure?
- 85°F
  - 50°C
  - 20 K
  - 545°F
- \_\_\_\_\_ 9. What is -175°C on the Kelvin scale?
- 76 K
  - 89 K
  - 98 K
  - 448 K
- \_\_\_\_\_ 10. The principles of atomic theory recognized today were conceived by
- Avogadro.
  - Bohr.
  - Dalton.
  - Rutherford.
- \_\_\_\_\_ 11. The total number of protons and neutrons in the nucleus of an atom is its
- atomic number.
  - Avogadro number.
  - mass number.
  - average atomic mass.
- \_\_\_\_\_ 12. All atoms of the same element have the same
- atomic mass.
  - number of neutrons.
  - mass number.
  - atomic number.
- \_\_\_\_\_ 13. Atoms of the same element can differ in
- chemical properties.
  - mass number.
  - atomic number.
  - number of protons and electrons.
- \_\_\_\_\_ 14. Chlorine has atomic number 17 and mass number 35. It has
- 17 protons, 17 electrons, and 18 neutrons.
  - 35 protons, 35 electrons, and 17 neutrons.
  - 17 protons, 17 electrons, and 52 neutrons.
  - 18 protons, 18 electrons, and 17 neutrons.
- \_\_\_\_\_ 15. Phosphorus-33 (atomic number 15) contains
- 33 protons.
  - 18 neutrons.
  - 33 neutrons.
  - 18 protons.
- \_\_\_\_\_ 16. Neon-22 contains 12 neutrons. It also contains
- 12 protons.
  - 22 protons.
  - 22 electrons.
  - 10 protons.
- \_\_\_\_\_ 17. Valence electrons determine an atom's

# Chemistry Final Exam Study Guide Multiple Choice

**EW Minium**



## **Chemistry Final Exam Study Guide Multiple Choice:**

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*Resources in Education* , 1983-04

*Study Guide for Introductory Chemistry* Iris Stovall, 1990

**Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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