

Chemistry 338 2nd Semester Final Exam Review Sheet

District Assessment

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten	¹⁸⁴ W ⁺² 74	74	74	110	72	184	+2

- Name the compound
 a. Fe_2O_3 – **iron (III) oxide**
 b. Cl_2O_8 – **dichlorine octaoxide**
- Write the formula
 a. Trisulfur Dinitride - **S_3N_2**
 b. Zinc Hydroxide – **$\text{Zn}(\text{OH})_2$**
 c. Tin (IV) oxide - **SnO_2**
- The pressure of a gas is 843 mm Hg at 24°C. What will the pressure be at standard temperature? Volume is constant.

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \frac{843\text{mmHg}}{297\text{K}} = \frac{P_2}{273\text{K}} \quad P_2 = 775 \text{ mm Hg}$$
- The volume of a gas is 81.7 mL when the pressure is 119 kPa and the temperature is 73°C. What is the volume at STP?

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2} \quad \frac{(119\text{kPa})(81.7\text{mL})}{346\text{K}} = \frac{(101.3\text{kPa})(V_2)}{273\text{K}} \quad V_2 = 75.7 \text{ mL}$$
- Between Be and Ba, which has a larger ionization energy? Larger atomic radius?
Larger ionization energy – Be
Larger atomic radius - Ba
- Balance the equation & classify the type of reaction:

$$2\text{Al} + 3 \text{CuSO}_4 \rightarrow 1\text{Al}_2(\text{SO}_4)_3 + 3 \text{Cu} \quad \text{single replacement}$$
- Predict the products & classify the type of reaction:

$$\text{H}_3\text{PO}_4 + 3 \text{NaOH} \rightarrow 3 \text{H}_2\text{O} + \text{Na}_3\text{PO}_4 \quad \text{double replacement}$$
- Calculate the Molar Mass: $(\text{NH}_4)_2\text{CO}_3$
96 g/mole
- Convert 2.75 moles of N_2O_5 to molecules

$$\frac{2.75\text{moles}}{1} \times \frac{6.02 \times 10^{23}\text{molecules}}{1\text{mole}} = 1.66 \times 10^{24} \text{ molecules } \text{N}_2\text{O}_5$$

Solutions

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