

## Gases | Study Guide

## Test Format

## Multiple Choice

## Gas Law Problems

## Gas Stoich Problems

## Short Answer

## Equations &amp; Conversions

$$P = \frac{F}{A}$$

$$1 \text{ atm} = 760 \text{ mmHg} = 101.3 \text{ kPa}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\frac{\text{Rate of Effusion of L}}{\text{Rate of Effusion of H}} = \frac{\sqrt{\text{Molar Mass of H}}}{\sqrt{\text{Molar Mass of L}}}$$

$$PV = nRT$$

$$K = {}^{\circ}\text{C} + 273 \quad 0.001 \text{ kL} = 1 \text{ L} = 1,000 \text{ mL}$$

$$R = 0.0821 \text{ L} \cdot \text{atm}/\text{mol} \cdot \text{K}$$

$$\text{At STP, 1 mole} = 22.4 \text{ L} \quad \text{KHDSdcm}$$

1. A box weighs 180 lbs and has a base that is 6 inches by 9 inches. What is the pressure this box exerts on the surface it is sitting on?

$$P = \frac{F}{A} \quad \frac{180 \text{ lbs}}{54 \text{ in}^2} = 3.3 \text{ psi}$$

$54 \text{ in}^2 = \text{area}$

2. What is atmospheric pressure?

The pressure exerted by the collisions of gases in the air surrounding Earth (N<sub>2</sub>, O<sub>2</sub>, etc.).

3. Convert 1.088 atm to:

a. mm Hg  $\frac{1.088 \text{ atm}}{1} \cdot \frac{760 \text{ mmHg}}{1 \text{ atm}} = 826.9 \text{ mmHg}$

b. kPa  $\frac{1.088 \text{ atm}}{1} \cdot \frac{101.3 \text{ kPa}}{1 \text{ atm}} = 110.2 \text{ kPa}$

4. What is a barometer and who invented it?

A tool that measures room/air pressure. Invented by Torricelli.



5. How are temperature and kinetic energy related?

Direct: As temp  $\uparrow$ , KE  $\uparrow$ , molecules move faster.  
As temp  $\downarrow$ , KE  $\downarrow$ , molecules move slower.

6. Briefly describe the five assumptions of the kinetic-molecular theory.

Gases have a lot of tiny particles that are always in motion that is random, fast, and continuous. Particles are far apart without any forces of attraction, so their collisions are elastic  $\rightarrow$  bouncing off of each other/walls. The speed of the molecules change with temperature.

# Chemistry Gases Study Guide Answers Teacher Guide

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In a digitally-driven world where displays reign supreme and quick communication drowns out the subtleties of language, the profound techniques and emotional nuances hidden within phrases usually move unheard. Yet, situated within the pages of **Chemistry Gases Study Guide Answers Teacher Guide** a interesting literary value blinking with raw feelings, lies an exceptional journey waiting to be undertaken. Written by an experienced wordsmith, this charming opus encourages readers on an introspective trip, delicately unraveling the veiled truths and profound impact resonating within ab muscles material of every word. Within the psychological depths of this poignant review, we can embark upon a sincere exploration of the book is primary themes, dissect their interesting writing model, and succumb to the effective resonance it evokes serious within the recesses of readers hearts.

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