

Chapter 2 Chemistry of Life

- Matter anything that takes up space
- An Element is one of the basic building blocks of matter and cannot be broken down by chemical means.
- It is even more surprising that that over 90% of the human body is composed of just four elements.
 - Carbon
 - Nitrogen
 - Oxygen
 - Hydrogen
- Atoms are the simplest unit of element that still retains the chemical and physical properties of an element.
 - There are the best-known subatomic particles: one positively charged protons
 - Uncharged: Neutrons
 - Negatively charged: Electrons
 - Protons and Neutrons are located within the nucleus of an atom and electrons move about the nucleus
 - The circle around the nucleus of an atom represents an electron shell, which represents the average location of electrons
- The atomic number represents the number of protons in the nucleus of an element
- Mass number of an atom represents the sum of protons and neutrons in the nucleus
- The atomic mass (the number below the atomic symbol on the periodic table) is the average of the AMU for all the isotopes of that atom. To determine the number of neutrons, subtract the number of protons from the atomic mass and take the closest whole number.
- Isotopes of an atom have the same number of protons but different numbers
- Atom with more than one shell are most stable when the outer shell is also called shell contains 8 electrons.
- Ions are particles that either carry either positive or negative charge.
- The attraction between oppositely charged sodium ions and chloride ions form an ionic bond.
- Atoms share in covalent bonds
- A hydrogen bond is the attraction of a slightly positive covalently bonded hydrogen to a slightly negative atom in the vicinity
- Water has a high heat capacity
- Water has a high heat of evaporation
- Water molecules are cohesive and adhesive
- Cohesion refers to the ability of water molecules to cling to each other
- Adhesion refers to the ability of water molecule to cling to a surface
- Lemon juice, Vinegar, Tomatoes, and coffee are all acidic solutions
- Acids are substances that dissociates in water releasing hydrogen ions. The acidity of a substance depends on how fully it dissociates in water

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Carl S. Pike

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Heitz,Jane B. Reece,2004 Table of contents continued How are water and good transported in plants What do you need to consider in order to grow plants in space or anywhere else for that matter How can plant reproduction be modified using biotechnology How do gravity and light affect plant growth responses How does an organism s structure help it maintain homeostasis How are form and function related in the digestive system How is mammalian heart structure related to function How do we breathe and why do we breathe How does the immune system keep the body free of pathogens What is nitrogenous waste and how is it removed from the body How do hormones regulate cell functions How does the production of male and female gametes differ in humans What common events occur in the early development of animals How do neurons function to transmit information What would happen if you modified a particular aspect of neuron function How does sarcomere structure affect muscle function What would happen if you modified particular aspects of muscle function What factors determine climate What determines behavior What methods can you use to determine population density and distribution What models can you use to calculate how quickly a population can grow What do you need to consider when analyzing communities of organisms What limits do available solar radiation and nutrients place on carrying capacities What factors can affect the survival of a species or community The activities of this workbook focus on key ideas principles and concepts that are basic to understanding biology The overall organization follows that of Campbell Reece Biology 7th edition

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