



## Reaction Rates and Equilibrium Study Guide

### Topics Covered:

Factors affecting reaction rates

Collision theory of reaction rates

Reaction progress diagrams (activation energy and transition states)

Rate laws and the method of initial rates

Chemical equilibrium

Equilibrium expressions and constants and their meaning

Product-favored vs. Reactant-favored

Shifting equilibrium and Le Chatelier's Principle

Quantitative calculations of equilibrium systems using ICE tables

### Practice Questions: [Click here for an answer key](#)

1. Describe how each of the following factors would affect the rate of "particle collisions" in your response.
  - a. Temperature decrease
  - b. Addition of iodide ions,  $I^-$ , as a catalyst
  - c. Increased reactant concentration
  - d. Crushing a solid into a powder (with larger surface area)
2. The reaction progress diagram shows the energy changes that take nitrogen dioxide gas.



# Chemistry Reaction Rates Study Guide

**Seager/Slabaugh**

## **Chemistry Reaction Rates Study Guide:**

**Chemical Principles Study Guide/Solutions Manual** John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

**Organic Chemistry Reactions: A Study Guide** Cybellium, 2024-10-26

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**Chemical Principles Student's Study Guide & Solutions Manual** John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

**Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry** Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section's critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

**Study Guide for Chang's Chemistry** Kenneth W. Watkins, 1988

**Chemistry, Study Guide** Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460

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Beiser,1971      Study Guide to Accompany Basics for Chemistry Martha Mackin,1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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