

18) What are the two products that are **always** formed in a combustion reaction?

_____ + _____

19) When balancing a chemical reaction, what are the **large numbers** that we adjust? _____

20) When balancing a chemical reaction, what are the **small numbers** in molecules? _____

21) What does the **Law of Conservation of Mass** tell us about both sides of a reaction? Explain.

In the questions below, put a **checkmark** next to the equations that are **not** balanced correctly. ————



Use the following reaction to answer the questions below:



30) In the reaction above, what does the **(aq)** stand for next to H_2O_2 ? _____

31) In the reaction above, what does the **(l)** stand for next to H_2O ? _____

32) In the reaction above, what does the **(g)** stand for next to O_2 ? _____

33) Knowing this information above, what would an **(s)** symbol stand for? _____

34) In the reaction above, list the **reactant(s)**: _____, the **product(s)**: _____

35) What is the arrow called? _____

36) In the expression **3 NO₂**, which number is the **coefficient** and which number is the **subscript**?

3 = _____ 2 = _____

Chemistry Reactions Study Guide

William M. Ritchey



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Table of Contents Chemistry Reactions Study Guide

1. Understanding the eBook Chemistry Reactions Study Guide
 - The Rise of Digital Reading Chemistry Reactions Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Reactions Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Reactions Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Reactions Study Guide

- Personalized Recommendations
 - Chemistry Reactions Study Guide User Reviews and Ratings
 - Chemistry Reactions Study Guide and Bestseller Lists
5. Accessing Chemistry Reactions Study Guide Free and Paid eBooks
 - Chemistry Reactions Study Guide Public Domain eBooks
 - Chemistry Reactions Study Guide eBook Subscription Services
 - Chemistry Reactions Study Guide Budget-Friendly Options
 6. Navigating Chemistry Reactions Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Reactions Study Guide Compatibility with Devices
 - Chemistry Reactions Study Guide Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Reactions Study Guide
 - Highlighting and Note-Taking Chemistry Reactions Study Guide
 - Interactive Elements Chemistry Reactions Study Guide
 8. Staying Engaged with Chemistry Reactions Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Reactions Study Guide
 9. Balancing eBooks and Physical Books Chemistry Reactions Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Reactions Study Guide
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Reactions Study Guide
 - Setting Reading Goals Chemistry Reactions Study Guide
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Reactions Study Guide

- Fact-Checking eBook Content of Chemistry Reactions Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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