

2-8) What are the two products that are always formed in a combustion reaction? _____

2-9) When balancing a chemical reaction, what are the large numbers that we adjust? _____

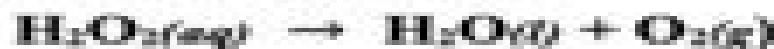
2-10) When balancing a chemical reaction, what are the small numbers in molecules? _____

2-11) What does the Law of Conservation of Mass tell us about both sides of a reaction? (Explain) _____

In the questions below, put a checkmark next to the responses that are just balanced correctly. _____



Use the following reaction to answer the questions below:



3-1) In the reaction above, what does the (l) stand for next to H_(l)? _____

3-2) In the reaction above, what does the (g) stand for next to O_(g)? _____

3-3) In the reaction above, what does the (aq) stand for next to H₂O_(aq)? _____

3-4) Knowing this information above, what would an (s) symbol stand for? _____

3-5) In the reaction above, list the reactant(s): _____ the product(s): _____

3-6) What is the arrow called? _____

3-7) In the expression 3 N₂O₂, which number is the coefficient and which number is the subscript? 3 = _____ 2 = _____

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William M. Ritchey

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Table of Contents Chemistry Reactions Study Guide

1. Understanding the eBook Chemistry Reactions Study Guide
 - The Rise of Digital Reading Chemistry Reactions Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Reactions Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Reactions Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Reactions Study Guide

- Personalized Recommendations
- Chemistry Reactions Study Guide User Reviews and Ratings
- Chemistry Reactions Study Guide and Bestseller Lists

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- Chemistry Reactions Study Guide Public Domain eBooks
- Chemistry Reactions Study Guide eBook Subscription Services
- Chemistry Reactions Study Guide Budget-Friendly Options

6. Navigating Chemistry Reactions Study Guide eBook Formats

- ePUB, PDF, MOBI, and More
- Chemistry Reactions Study Guide Compatibility with Devices
- Chemistry Reactions Study Guide Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Chemistry Reactions Study Guide
- Highlighting and Note-Taking Chemistry Reactions Study Guide
- Interactive Elements Chemistry Reactions Study Guide

8. Staying Engaged with Chemistry Reactions Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Reactions Study Guide

9. Balancing eBooks and Physical Books Chemistry Reactions Study Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chemistry Reactions Study Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chemistry Reactions Study Guide

- Setting Reading Goals Chemistry Reactions Study Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Reactions Study Guide

- Fact-Checking eBook Content of Chemistry Reactions Study Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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