

Chemistry Review Packet

Name: _____

Date: _____

1. Consider the following reaction:



When 2.0 moles of hydrogen are reacted with 48.0 grams of oxygen, how many grams of water are formed?

- A. 16.0 grams B. 26.0 grams C. 34.0 grams D. 44.0 grams E. 50.0 grams

2. Which of the following 1.0 M aqueous solutions has the lowest pH?

- A. HNO_3 B. Na_2CO_3 C. HNO_2 D. NaOH E. HCO_2H

3. Which of the following substances cannot be broken down into simpler substances by physical or chemical means?

- A. Silver B. Baking soda C. Natural gas D. Neon gas E. Water

4. What is the approximate percent composition by mass of hydrogen in the compound ethanol, $\text{C}_2\text{H}_6\text{O}$ (Molecular weight, in grams: 46.12 to 46.13; 46.00 to 46.02)

- A. 4.3% B. 12.0% C. 36.0% D. 54.0% E. 20.0%

5. The empirical formula of a certain hydrocarbon is C_2H_5 , and it has a molecular weight equal to 60 g/mol. Based on this information, the molecular formula of the hydrocarbon is

- A. C_2H_5 B. C_4H_{10} C. C_6H_{15} D. C_8H_{20} E. $\text{C}_{10}\text{H}_{25}$

6. Which of the following aqueous solutions at 25°C has the highest normal boiling point?

- A. 0.20 M CaCl_2 B. 0.20 M NaCl C. 0.5 M KNO_3 D. 1.0 M $\text{C}_6\text{H}_{12}\text{O}_6$ E. 0.20 M MgCl_2

7. Consider the reaction



How is this reaction best classified?

- A. Precipitation B. Redox reaction C. Single replacement D. Combustion E. Synthesis

8. The following reaction takes place under constant temperature and pressure:



When 10.0 L of hydrogen gas reacts completely with excess chlorine, how many liters of gas remain? Assume

- A. 1.00 L B. 1.50 L C. 10.0 L D. 15.0 L E. 20.0 L

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