

Beispiel:

Stärke 0,001 mol/l

① Lösung:

$C_{H_2O} = 0,01 \text{ mol/l}$

$K_1 = 3,2 \cdot 10^{-7}$

$\text{pH} = ?$

Reaktion:

Der ersten Reaktion:

$[\text{H}^+] = \sqrt{K_1 \cdot C_{\text{H}_2\text{O}}}$

$= \sqrt{3,2 \cdot 10^{-7} \cdot 0,01} = 2,8 \cdot 10^{-5} \text{ mol/l}$

$\text{pH} = -\lg [\text{H}^+] = -\lg 2,8 \cdot 10^{-5} = 4,78 + 0,0057 = \underline{\underline{4,78}}$

② Lösung:

$\text{pH}_1 = 6,3 \cdot 10^6$

$\text{pH} = ?$

Reaktion:

$\text{Ca(OH)}_2 = \text{Ca}^{2+} + 2\text{OH}^-$

$\text{pH} = [\text{Ca}^{2+}] [\text{OH}^-]^2 = 6 \cdot (2,5)^2 = 45$

S - hydrolyse nach:

$S = \sqrt{\frac{\text{pH}}{4}} = \sqrt{\frac{6,3 \cdot 10^6}{4}} = 1,62 \cdot 10^{-3} \text{ mol/l}$

$[\text{OH}^-] = \frac{S}{2} = \frac{1,62 \cdot 10^{-3}}{2} = 8 \cdot 10^{-4} \text{ mol/l}$

$[\text{H}^+] = \frac{10^{-14}}{8 \cdot 10^{-4}} = \frac{10^{-14}}{8 \cdot 10^{-3}} = 1,25 \cdot 10^{-12}$

$\text{pH} = -\lg [\text{H}^+] = -\lg 1,25 \cdot 10^{-12} = 12 - 0,925 = \underline{\underline{11,07}}$

③ Lösung:

Volumen: 1000 ml

$C_{\text{H}_2\text{O}} = 0,1 \text{ mol/l}$

$\text{Volumen: } 1000 \text{ ml}$

$C_{\text{Ca(OH)}_2} = 0,01 \text{ mol/l}$

$\text{pH}_1 = 6,6 \cdot 10^6$

Reaktion:

Reaktion:

$C_{\text{Ca}^{2+}} = C_{\text{OH}^-} = 0,01 \text{ mol/l}$

$C_{\text{Ca}^{2+}} = C_{\text{OH}^-} = 0,01 \text{ mol/l}$

$\text{pH}_1 = \text{pH}_2 = 6,6 \cdot 10^6$

$\text{pH} = \frac{\text{pH}_1 + \text{pH}_2}{2} = \frac{6,6 \cdot 10^6 + 6,6 \cdot 10^6}{2} = 6,6 \cdot 10^6 \text{ mol/l}$

$[\text{H}^+] [\text{OH}^-]^2 = 10^{-14}$

$[\text{H}^+]^2 \cdot 10^{-14} = 10^{-14}$

$[\text{H}^+]^2 = 10^{-14}$

$[\text{H}^+] = \sqrt{10^{-14}} = 10^{-7} \text{ mol/l}$

$\text{pH} = -\lg [\text{H}^+] = -\lg 10^{-7} = 7$

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web may 29 2018 figure 2 the graph shows the solubility curve of potassium nitrate with corresponding temperatures as the temperature increases so does the solubility of the solution a y 0 2 1 61 b y 0 2 1 61 0 60 2 1 60 61 0 40 2 1 40 61 8 5 g 100 ml 4 6 g 100 ml

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web mar 15 2021 this experiment will further develop your understanding of thermodynamics while relating it to two concepts you already have studied equilibrium and solubility you will determine three thermodynamic values associated with the solubility of potassium nitrate the reaction we will be studying is the dissolution of potassium nitrate in water

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web analysis construct a solubility curve using the mass of solute that will dissolve in 100 g of water at a specified temperature in this lab 5 ml of water was used assume 1 0 g ml of water take the collected data make it proportional to get the solubility of the kno3 graph temperature vs solubility using a best fit curve

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web part i construct a single graph with multiple solubility curves to clearly show the solubility of the four ionic compounds in the data table below the following tips will help you make the best possible graph