

1. How many atoms are in 4 moles of iron?

$$4 \times 6.02 \times 10^{23} = 2.4 \times 10^{24}$$

2. What is the percentage by mass of oxygen in water? H_2O

$$\frac{16}{18} \times 100 = 89\%$$

$$H_2 = 2 \times 1 = 2$$

$$O = 1 \times 16 = \frac{16}{18}$$

3. A compound is 57.5 % sodium, 40% oxygen and 2.5 % hydrogen. What is its empirical formula?

$$57.5g \times \frac{1\text{mol}}{23g\text{ Na}} = 2.5\text{ mol Na}$$

$$2.5\text{ g H} \times \frac{1\text{mol H}}{1g\text{ H}} = 2.5\text{ mol H}$$

$$40g \times \frac{1\text{mol O}}{16g\text{ O}} = 2.5\text{ mol O}$$

$$1:1:1 \text{ ratio}$$



4. A compound with an empirical formula of CH_6 has a molecular weight of 42. What is its molecular formula?



$$C = 3 \times 12 = 36$$

$$H = 6 \times 1 = \frac{6}{42}$$

44.6

5. At STP ~~48~~ liters of neon would contain the same number of molecules as how many grams of Argon? What volume would the argon occupy?

$$48\text{ L Ne} \times \frac{1\text{mol Ne}}{22.4\text{ L Ne}} = 2\text{ mol Ne}$$

$$2\text{ mol Ar} \times \frac{39.95\text{ g}}{1\text{mol Ar}} = 79.9\text{ g Ar}$$

44.8

VOLUME of Ar = 44.8 L

6. Sodium and chlorine react to form sodium chloride. If 4 moles of chlorine react, how many moles of sodium chloride will form?



$$4\text{ mol Cl}_2 \times \frac{2\text{ mol Na}}{1\text{ mol Cl}_2} = 8\text{ mol NaCl}$$

7. How many liters of chlorine at STP would be needed to react with 92 grams of sodium?

$$92\text{ g Na} \times \frac{1\text{mol Na}}{23\text{ g Na}} \times \frac{1\text{ mol Cl}_2}{2\text{ mol Na}} \times \frac{22.4\text{ L}}{1\text{ mol Cl}_2} = 44.8\text{ L Cl}_2$$

8. How many milliliters of chlorine were required in question 7?

$$44.8\text{ L} \times \frac{1000\text{ mL}}{1\text{ L}} = 44,800\text{ mL Cl}_2$$

9. Write and balance the equation for the combustion of butane.



10. If 232.0 grams of butane are combined with 192 grams of oxygen and a reaction occurs, how many grams of water will form? $232\text{ g C}_4\text{H}_{10} \times \frac{1\text{ mol C}_4\text{H}_{10}}{52\text{ g C}_4\text{H}_{10}} \times \frac{10\text{ mol H}_2\text{O}}{2\text{ mol C}_4\text{H}_{10}} \times \frac{16\text{ g H}_2\text{O}}{1\text{ mol H}_2\text{O}} = 360\text{ g H}_2\text{O}$

$$192\text{ g O}_2 \times \frac{1\text{ mol O}_2}{32\text{ g O}_2} \times \frac{10\text{ mol H}_2\text{O}}{13\text{ mol O}_2} \times \frac{16\text{ g H}_2\text{O}}{1\text{ mol H}_2\text{O}} = 83.9\text{ g H}_2\text{O}$$

11. If the reaction in number 9 was performed and 2 liters of carbon dioxide were produced at STP, what was the percent yield? $2\text{ L CO}_2 \times \frac{1\text{ mol CO}_2}{22.4\text{ L CO}_2} \times \frac{16\text{ g H}_2\text{O}}{8\text{ mol CO}_2} \times \frac{16\text{ g H}_2\text{O}}{1\text{ mol H}_2\text{O}} = 2\text{ g H}_2\text{O}$

$$\frac{2\text{ g H}_2\text{O}}{83.9\text{ g H}_2\text{O}} \times 100 = \boxed{2.4\%}$$

12. Write and balance the equation for the reaction of aluminum with hydrochloric acid.



$$\frac{2\text{ mol Al}}{26.98\text{ g Al}} \times \frac{2\text{ mol HCl}}{1\text{ mol Al}} \times \frac{6\text{ mol HCl}}{2\text{ mol Al}} \times \frac{34.5\text{ g HCl}}{1\text{ mol HCl}} = 146\text{ g HCl}$$

13. If you are given 108 grams of aluminum to react with 146 grams of HCl, what will be the limiting reactant?

$$108\text{ g Al} \times \frac{1\text{ mol Al}}{26.98\text{ g Al}} \times \frac{3\text{ mol HCl}}{1\text{ mol Al}} \times \frac{2\text{ g HCl}}{1\text{ mol HCl}} = 12\text{ g HCl}$$

$$146\text{ g HCl} \times \frac{1\text{ mol HCl}}{34.5\text{ g HCl}} \times \frac{3\text{ mol Al}}{1\text{ mol HCl}} \times \frac{2\text{ g Al}}{1\text{ mol Al}} = \boxed{4\text{ g Al}}$$

HCl is limiting reactant

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1. Understanding the eBook Chemistry Semester 2 Practice Final
 - The Rise of Digital Reading Chemistry Semester 2 Practice Final
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Semester 2 Practice Final
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Semester 2 Practice Final
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Semester 2 Practice Final
 - Personalized Recommendations
 - Chemistry Semester 2 Practice Final User Reviews and Ratings
 - Chemistry Semester 2 Practice Final and Bestseller Lists

5. Accessing Chemistry Semester 2 Practice Final Free and Paid eBooks
 - Chemistry Semester 2 Practice Final Public Domain eBooks
 - Chemistry Semester 2 Practice Final eBook Subscription Services
 - Chemistry Semester 2 Practice Final Budget-Friendly Options
6. Navigating Chemistry Semester 2 Practice Final eBook Formats
 - ePUB, PDF, MOBI, and More
 - Chemistry Semester 2 Practice Final Compatibility with Devices
 - Chemistry Semester 2 Practice Final Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Semester 2 Practice Final
 - Highlighting and Note-Taking Chemistry Semester 2 Practice Final
 - Interactive Elements Chemistry Semester 2 Practice Final
8. Staying Engaged with Chemistry Semester 2 Practice Final
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Semester 2 Practice Final
9. Balancing eBooks and Physical Books Chemistry Semester 2 Practice Final
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Semester 2 Practice Final
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Semester 2 Practice Final
 - Setting Reading Goals Chemistry Semester 2 Practice Final
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Semester 2 Practice Final
 - Fact-Checking eBook Content of Chemistry Semester 2 Practice Final
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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