

1. How many atoms are in 4 moles of iron?

$$4 \times 6.02 \times 10^{23} = 2.4 \times 10^{24}$$

2. What is the percentage by mass of oxygen in water?

$$\frac{16}{18} \times 100 = 89\%$$



$$H_2 = 2 \times 1 = 2$$

$$O = 1 \times 16 = 16$$

3. A compound is 57.5 % sodium, 40% oxygen and 2.5 % hydrogen. What is its empirical formula?

$$57.5g \times \frac{1 \text{ mol}}{23g \text{ Na}} = 2.5 \text{ mol Na}$$

$$2.5g \text{ H} \times \frac{1 \text{ mol}}{1g \text{ H}} = 2.5 \text{ mol H}$$

$$40g \times \frac{1 \text{ mol}}{16g \text{ O}} = 2.5 \text{ mol O}$$

$$1:1:1 \text{ ratio}$$



4. A compound with an empirical formula of  $CH_2$  has a molecular weight of 42. What is its molecular formula?



$$C = 3 \times 12 = 36$$

$$H = 6 \times 1 = 6$$

$$42$$

5. At STP ~~48~~ 44.8 liters of neon would contain the same number of molecules as how many grams of Argon? What volume would the argon occupy?

$$44.8 \text{ L Ne} \times \frac{1 \text{ mol Ne}}{22.4 \text{ L Ne}} = 2 \text{ mol Ne}$$

$$2 \text{ mol Ar} \times \frac{39.9g \text{ Ar}}{1 \text{ mol Ar}} = 79.8g \text{ Ar}$$

$$\text{volume of Ar} = 44.8 \text{ L}$$

6. Sodium and chlorine react to form sodium chloride. If 4 moles of chlorine react, how many moles of sodium chloride will form?



$$4 \text{ mol } Cl_2 \times \frac{2 \text{ mol Na}}{1 \text{ mol } Cl_2} = 8 \text{ mol Na}$$

7. How many liters of chlorine at STP would be needed to react with 92 grams of sodium?

$$92g \text{ Na} \times \frac{1 \text{ mol Na}}{23g \text{ Na}} \times \frac{1 \text{ mol } Cl_2}{2 \text{ mol Na}} \times \frac{22.4 \text{ L}}{1 \text{ mol } Cl_2} = 44.8 \text{ L } Cl_2$$

8. How many milliliters of chlorine were required in question 7?

$$44.8 \text{ L} \times \frac{1000 \text{ mL}}{1 \text{ L}} = 44,800 \text{ mL } Cl_2$$

9. Write and balance the equation for the combustion of butane.



10. If 232.0 grams of butane are combined with 192 grams of oxygen and a reaction occurs, how many grams of water will form?

$$232g \text{ C}_4\text{H}_{10} \times \frac{1 \text{ mol C}_4\text{H}_{10}}{58g \text{ C}_4\text{H}_{10}} \times \frac{10 \text{ mol H}_2\text{O}}{2 \text{ mol C}_4\text{H}_{10}} \times \frac{18g \text{ H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 360g \text{ H}_2\text{O}$$

$$192g \text{ O}_2 \times \frac{1 \text{ mol O}_2}{32g \text{ O}_2} \times \frac{10 \text{ mol H}_2\text{O}}{13 \text{ mol O}_2} \times \frac{18g \text{ H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{83g \text{ H}_2\text{O}}$$

11. If the reaction in number 9 was performed and 2 liters of carbon dioxide were produced at STP, what was the percent yield?

$$2 \text{ L } CO_2 \times \frac{1 \text{ mol } CO_2}{22.4 \text{ L } CO_2} \times \frac{10 \text{ mol H}_2\text{O}}{8 \text{ mol } CO_2} \times \frac{18g \text{ H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 2g \text{ H}_2\text{O}$$

$$\frac{2g \text{ H}_2\text{O}}{83g \text{ H}_2\text{O}} \times 100 = \boxed{2.4\%}$$

12. Write and balance the equation for the reaction of aluminum with hydrochloric acid.



$$\frac{2g \text{ H}_2}{2g \text{ H}_2} \times \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} \times \frac{27g \text{ Al}}{1 \text{ mol}} = 13.5g \text{ Al}$$

13. If you are given 108 grams of aluminum to react with 146 grams of HCl, what will be the limiting reactant?

$$108g \text{ Al} \times \frac{1 \text{ mol Al}}{26.98g \text{ Al}} \times \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} \times \frac{2g \text{ H}_2}{1 \text{ mol H}_2} = 12g \text{ H}_2$$

$$146g \text{ HCl} \times \frac{1 \text{ mol HCl}}{36.45g \text{ HCl}} \times \frac{3 \text{ mol H}_2}{6 \text{ mol HCl}} \times \frac{2g \text{ H}_2}{1 \text{ mol H}_2} = \boxed{4g \text{ H}_2}$$

HCl is limiting reactant

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web josue ortega city state zip code cell 000 000 0000 email email com felicity weeks please let this letter serve as my official recommendation for belinda carlisle for the position of executive assistant

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