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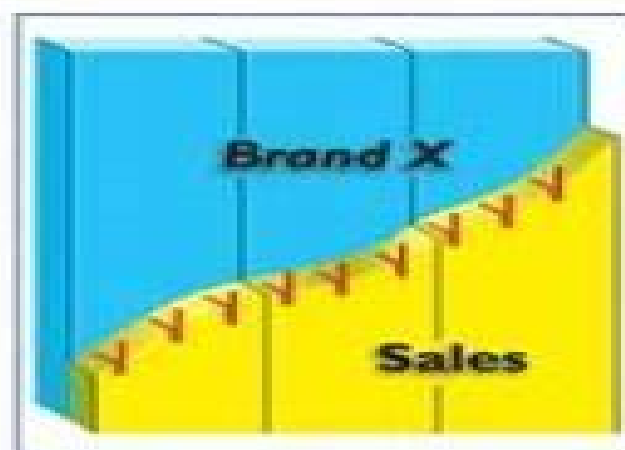
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web consider the thermochemical equation below $4\text{NO(g)} + 6\text{H}_2\text{O(l)} + 4\text{NH}_3\text{(g)} + 5\text{O}_2\text{(g)} \rightarrow \text{rh } 1170\text{ kJ}$ a is this reaction endothermic or exothermic endothermic rh gt 0 b how much heat will be absorbed or released if 10 g of no is reacted with 10 g h₂o

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web thermochemistry answers and explanations 1 a a thermodynamic system has two means of exchanging energy with its surroundings heat flow and pressure volume work heat flow q is the transfer of thermal energy due to temperature difference work w is macroscopic energy transfer between the system and its surroundings that occurs

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