

# CHEM 121: Final Exam Study Guide

## Chapter 1

- Know the Scientific method
- Know the definitions for hypothesis, scientific law, and scientific theory.

## Chapter 2

- Length, mass, weight, volume
  - Know  $1 \text{ cm}^3 = 1 \text{ mL}$  and  $1 \text{ dm}^3 = 1 \text{ L}$
- Significant figures and Scientific notation
  - Rounding, in addition or subtraction, in multiplication or division
  - Keep as many sig figs until final answer
- Scientific notation
- Metric system
  - Know the metric prefixes: kilo (k), deci (d), centi (c), milli (m), and micro (μ)
  - Be able to perform metric-metric conversions using these prefixes.
- Use the metric-English conversions provided (1 in. = 2.54 cm; 1 lb = 453.6 g; 1 qt = 946 mL)

- Volume by calculation
  - $V_{\text{rectangular\ solid}} = \text{length} \times \text{width} \times \text{thickness}$
- Volume by displacement
- Density:  $d = \frac{m}{V}$ 
  - Calculate density, mass, or volume
  - Identify what items sink or float given densities of liquids and solids.
- Temperature
  - Know the formulas for converting "F-to-"C or "C-to-"F and K-to-"C or "C-to-K
- Percentage: ratio of parts per 100 parts
  - Given amount of part and whole, calculate %
  - Use a given % to solve for part or whole

## Chapter 3

- Know that matter is studied at the macroscopic, microscopic, particulate (molecular) levels
- Physical states of matter
  - Determine physical state of substances (solids, liquids, gases) given descriptions of volume, shape, particles moving, etc.
- Classify properties and changes as physical or chemical
  - Know terms for changes of state:
    - Melting, freezing, vaporizing, condensation, sublimation, deposition

- Classification of matter
  - Given examples, determine which are elements, compounds, or mixtures
  - Given molecular-level images, determine which are elements, compounds, or mixtures and solids, liquids, or gases
  - Distinguish between homogeneous and heterogeneous mixtures
- Chemical reaction:
  - reactants: starting materials
  - products: substances produced in reaction

# Chem College Final Exam Study Guide

**M Woodhall**

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**PSI Study Guide for Organic Chemistry by Douglas C. Neckers, Michael P. Doyle** E. C. Blossey, 1977

**Student's Guide to Chemistry; a Modern Introduction** David Brooks, 1974      **Learning Science Through**

**Computer Games and Simulations** National Research Council, Division of Behavioral and Social Sciences and Education, Board on Science Education, Committee on Science Learning: Computer Games, Simulations, and Education, 2011-05-12 At a time when scientific and technological competence is vital to the nation's future the weak performance of U.S. students in science reflects the uneven quality of current science education. Although young children come to school with innate curiosity and intuitive ideas about the world around them science classes rarely tap this potential. Many experts have called for a new approach to science education based on recent and ongoing research on teaching and learning. In this approach simulations and games could play a significant role by addressing many goals and mechanisms for learning science: the motivation to learn science, conceptual understanding, science process skills, understanding of the nature of science, scientific discourse and argumentation, and identification with science and science learning. To explore this potential, *Learning Science Computer Games Simulations and Education* reviews the available research on learning science through interaction with digital simulations and games. It considers the potential of digital games and simulations to contribute to learning science in schools, in informal out-of-school settings, and everyday life. The book also identifies the areas in which more research and research-based development is needed to fully capitalize on this potential. *Learning Science* will guide academic researchers, developers, publishers, and entrepreneurs from the digital simulation and gaming community and education practitioners and policy makers toward the formation of research and development partnerships that will facilitate rich intellectual collaboration. Industry, government agencies, and foundations will play a significant role through start-up and ongoing support to ensure that digital games and simulations will not only excite and entertain but also motivate and educate.

*Resources in Education*, 1984      **McGraw-Hill's 500 College Chemistry Questions** David E. Goldberg, 2012-09-07 500

Ways to Achieve Your Best Grades We want you to succeed on your college chemistry midterm and final exams. That's why we've selected these 500 questions to help you study more effectively, use your preparation time wisely, and get your best grades. These questions are similar to the ones you'll find on a typical college exam, so you will know what to expect on test day. Each question includes comprehensive explanations in the answer key. Whether you have been studying all year or are doing a last-minute review, *McGraw-Hill's 500 College Chemistry Questions* will help you achieve the final grade you desire. Sharpen your subject knowledge and build your test-taking confidence with 500 essential college chemistry questions with answers. Clear solutions in the answer key for every problem. Coverage from atomic mass to electrochemistry.

*The Canada Medical Record*, 1892      **Florida Schools**, 1964      **Manual of Chemistry** William Simon, 1884      **Chemical News and Journal of Industrial Science**, 1907      **El-Hi Textbooks in Print**, 1984      *The Publishers' Trade List Annual*, 1976      **Technical**

**Education and Industrial Training** ,1966      **New York Herald Tribune Book Review** ,1959      **Paperbound Books in Print** ,1966      **Matriculation chemistry papers. The last 32 papers. [Continued as] The papers from Jan. 1890 to June 1907 (to Jan. 1908, to Sept. 1908, Jan. 1894 to Sept. 1909, to June 1910, to Jan. 1911, Jan. 1906 to June 1915).** London univ, exam. papers,1891      **Science chemistry papers. Being the papers for 1869-1889 (1869 to 1899, to 1900, to 1904, 1882 to 1905, to 1906, to 1907, 1892 to 1908, to 1909, to 1910, to 1911).** London univ, exam. papers,1889      **American Druggists' Circular and Chemical Gazette** ,1883      **Elementary Studies in Chemistry** Joseph Torrey (jr.),1899

**A Manual of Chemical Analysis as Applied to the Examination of Medicinal Chemicals** Friedrich Hoffmann, Frederick Belding Power,1883      *The National Guide to Educational Credit for Training Programs* American Council on Education,2005 Highlights over 6 000 educational programs offered by business labor unions schools training suppliers professional and voluntary associations and government agencies

## Embracing the Track of Phrase: An Psychological Symphony within **Chem College Final Exam Study Guide**

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