



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time having reached Equilibrium. Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time having reached Equilibrium. Equilibrium reactions are represented by the symbol \rightleftharpoons.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant</p> <p>The equilibrium constant is a measure of the extent to which a reaction proceeds. It is given by the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$									
<p>Equilibrium Constant, K_c</p> <p>The equilibrium constant is a measure of the extent to which a reaction proceeds. It is given by the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K_c</p> <table><tr><th>Equilibrium Constant, K_c</th><th>Meaning</th></tr><tr><td>$K_c > 1$</td><td>Concentration of products is greater than concentration of reactants.</td></tr><tr><td>$K_c < 1$</td><td>Concentration of reactants is greater than concentration of products.</td></tr><tr><td>$K_c = 1$</td><td>Concentration of reactants is equal to concentration of products.</td></tr></table>		Equilibrium Constant, K_c	Meaning	$K_c > 1$	Concentration of products is greater than concentration of reactants.	$K_c < 1$	Concentration of reactants is greater than concentration of products.	$K_c = 1$	Concentration of reactants is equal to concentration of products.
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<p>Equilibrium Constant, K_p</p> <p>The equilibrium constant is a measure of the extent to which a reaction proceeds. It is given by the ratio of the partial pressures of the products to the partial pressures of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	<p>Equilibrium Constant, K_p</p> <table><tr><th>Equilibrium Constant, K_p</th><th>Meaning</th></tr><tr><td>$K_p > 1$</td><td>Partial pressure of products is greater than partial pressure of reactants.</td></tr><tr><td>$K_p < 1$</td><td>Partial pressure of reactants is greater than partial pressure of products.</td></tr><tr><td>$K_p = 1$</td><td>Partial pressure of reactants is equal to partial pressure of products.</td></tr></table>		Equilibrium Constant, K_p	Meaning	$K_p > 1$	Partial pressure of products is greater than partial pressure of reactants.	$K_p < 1$	Partial pressure of reactants is greater than partial pressure of products.	$K_p = 1$	Partial pressure of reactants is equal to partial pressure of products.
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Le Chatelier's Principle: the effect of changing conditions on equilibrium										
Condition	Effect on Equilibrium	Example								
Concentration	<ul style="list-style-type: none">Increasing concentration of a reactant will shift the equilibrium to the right.Increasing concentration of a product will shift the equilibrium to the left.	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the concentration of N_2 will shift the equilibrium to the right, producing more NH_3.</p>								
Pressure	<ul style="list-style-type: none">Increasing pressure will shift the equilibrium to the side with the fewest moles of gas.Decreasing pressure will shift the equilibrium to the side with the most moles of gas.	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the pressure will shift the equilibrium to the right, producing more NH_3.</p>								
Temperature	<ul style="list-style-type: none">Increasing temperature will shift the equilibrium to the side that is endothermic.Decreasing temperature will shift the equilibrium to the side that is exothermic.	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the temperature will shift the equilibrium to the left, producing more N_2 and H_2.</p>								

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Victor M. Corman



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Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples Resources in Education, 1980 **Research in Education**, 1971 **Study Guide for Chemical Principles** Thomas Elliott Taylor, 1979 **Resources in Education**, 1977 **Chemistry: Matter & Change, Study Guide For Content Mastery, Student Edition** McGraw Hill, 2001-03-30 Study Guide and Reinforcement Worksheets allow for differentiated instruction through a wide range of question formats There are worksheets and study tools for each section of the text that help teachers track students progress toward understanding concepts Guided Reading Activities help students identify and comprehend the important information in each chapter **Chemistry**, 2001-04-01

Chemistry Steven S. Zumdahl, Susan A. Zumdahl, 2000 Zumdahl's conceptual approach to problem solving early and thorough treatment of reactions and integration of descriptive chemistry and chemical principles throughout are key elements to the success of this bestseller By encouraging students to understand qualitative concepts before considering quantitative problems Chemistry helps students build a solid understanding of topics and avoid common misconceptions The text also includes new Chemical Impact boxes that present everyday applications of chemistry plus over 200 redesigned illustrations with greater clarity and emphasis on molecular level illustrations Chemistry, Study Guide Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460 **Current Index to**

Journals in Education, 1991 **Chemical Reactions 2 : Equilibrium and Oxidation-reduction : CHE-5043-2 : Learning Guide** Chunn, Barbara, Dumas, André, Céline Tremblay, Moisan, Mireille, Société de formation à distance des commissions scolaires du Québec, Morazain, Paule, De Fulviis, Claudia, 2001 **Study Guide, Chemical Principles, Fifth Edition, Zumdahl** Paul B. Kelter, Steven S. Zumdahl, 2004-04 Conference Papers Index, 1982 Monthly Papers presented at recent meeting held all over the world by scientific technical engineering and medical groups Sources are meeting programs and abstract publications as well as questionnaires Arranged under 17 subject sections 7 of direct interest to the life scientist Full programs of meetings listed under sections Entry gives citation number paper title name mailing address and any ordering number assigned Quarterly and annual indexes to subjects authors and programs not available in monthly issues **The Spectator**, 1968 *Chemical Principles Student's Study Guide & Solutions Manual* John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text **Chemical Equilibrium In a Nutshell** Brian

Coleman,2018-06-25 Understanding the math and minutiae of chemical equilibrium can be a tall task for anyone so why not enlist the help of a scientific squirrel to guide you on your journey Join Dr Wash as we dabble in equilibrium constants and other tools needed to predict chemical processes This book focuses on introductory concepts at the high school and early university level focusing on identifying equilibrium calculating K and Q discussing Le Chatelier s principle and tying equilibrium with the field of Thermodynamics Full of step by step instructions and practice questions this book aims to simplify one of the more complex topics found within the field of chemistry Chemical Equilibrium William Guenther,1975-12 The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation Organic Chemistry Study Guide Robert J. Ouellette,J. David Rawn,2014-11-04 Organic Chemistry Study Guide Key Concepts Problems and Solutions features hundreds of problems from the companion book Organic Chemistry and includes solutions for every problem Key concept summaries reinforce critical material from the primary book and enhance mastery of this complex subject Organic chemistry is a constantly evolving field that has great relevance for all scientists not just chemists For chemical engineers understanding the properties of organic molecules and how reactions occur is critically important to understanding the processes in an industrial plant For biologists and health professionals it is essential because nearly all of biochemistry springs from organic chemistry Additionally all scientists can benefit from improved critical thinking and problem solving skills that are developed from the study of organic chemistry Organic chemistry like any skill is best learned by doing It is difficult to learn by rote memorization and true understanding comes only from concentrated reading and working as many problems as possible In fact problem sets are the best way to ensure that concepts are not only well understood but can also be applied to real world problems in the work place Helps readers learn to categorize analyze and solve organic chemistry problems at all levels of difficulty Hundreds of fully worked practice problems all with solutions Key

concept summaries for every chapter reinforces core content from the companion book **Chemistry Equations & Answers** Speedy Publishing, 2014-08-26 General chemistry inorganic chemistry organic chemistry and biochemistry are all difficult courses requiring much memorization for the student Essentially there is no easy way to learn formulas and facts This is why chemistry classes are such challenges to students even the best ones However a chemistry equations and answers study guide can help the student When used as a quick reference guide it can be used often to determine the formulas needed for various questions The astute student can cleverly devise ways to make the guide useful for test questions or other circumstances requiring one of the many chemistry equations *Chemical principles* Paul B. Kelter, 1998

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