

## 16 • Chemical Equilibrium

1. Consider the reaction system.  
 $\text{CoO(s)} + \text{H}_2\text{g} \rightleftharpoons \text{Co(s)} + \text{H}_2\text{O(g)}$ .

The equilibrium constant expression is

a)  $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$       d)  $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$   
 b)  $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$       c)  $\frac{[\text{H}_2\text{O}]}{[\text{H}_2]}$   
 (d)  $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$   
 (e)  $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$

2. Given the equilibrium,  
 $2\text{SO}_2\text{g} + \text{O}_2\text{g} \rightleftharpoons 2\text{SO}_3\text{g}$ , if this equilibrium is established by beginning with equal number of moles of  $\text{SO}_2$  and  $\text{O}_2$  in a 1.0 Liter bulb, then the following *must* be true at equilibrium:  
 a)  $[\text{SO}_2] = [\text{SO}_3]$       d)  $[\text{SO}_2] < [\text{O}_2]$   
 (d)  $2[\text{SO}_2] = 2[\text{SO}_3]$       e)  $[\text{SO}_2] > [\text{O}_2]$   
 c)  $[\text{SO}_2] = [\text{O}_2]$

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole  $\text{NO}$ , 0.200 mol  $\text{Cl}_2$  and 0.500 mol  $\text{ClNO}$  were placed in a 25.0 Liter container. The following equilibrium is established:  $2\text{ClNO(g)} \rightleftharpoons 2\text{NO(g)} + \text{Cl}_2\text{g}$

3. At equilibrium, 0.600 mol of  $\text{ClNO}$  was present. The number of *moles* of  $\text{Cl}_2$  present at equilibrium is  
 (a) 0.050      (d) 0.200  
 b) 0.100      e) 0.250  
 c) 0.150

## PRACTICE TEST

4. The equilibrium constant,  $K_c$ , is:  
 a)  $4.45 \times 10^{-4}$       d) 0.167  
 b)  $6.67 \times 10^{-4}$       e) 1500  
 (c) 0.111

5. At  $985^\circ\text{C}$ , the equilibrium constant for the reaction,  
 $\text{H}_2\text{g} + \text{CO}_2\text{g} \rightleftharpoons \text{H}_2\text{O(g)} + \text{CO(g)}$

is 1.63. What is the equilibrium constant for the reverse reaction?  
 a) 1.63      d) 0.613  
 b) 0.815      (c) 1.00  
 c) 2.66

6. What is the relationship between  $K_p$  and  $K_c$  for the reaction,  $2\text{ICl(g)} \rightleftharpoons \text{I}_2\text{g} + \text{Cl}_2\text{g}$ ?  
 (a)  $K_p = K_c(RT)^{-1}$       d)  $K_p = K_c$   
 b)  $K_p = K_c(RT)$       e)  $K_p = K_c(2RT)$   
 c)  $K_p = K_c(RT)^2$

7. For the reaction  $2\text{NO}_2\text{g} \rightleftharpoons \text{N}_2\text{O}_4\text{g}$ ,  $K_p$  at  $25^\circ\text{C}$  is 7.3, when all partial pressures are expressed in atmospheres. What is  $K_c$  for this reaction? [ $R=0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$ ]  
 (a) 4270      d) 179  
 b) 0.0119      e) 2.06  
 c) 0.291

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**Frank Brescia**

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**Fundamentals of Chemistry Laboratory Studies** Frank Brescia, 2012-12-02 Fundamentals of Chemistry Laboratory Studies focuses on the techniques involved in chemical laboratory operations Divided into 13 parts the manual gives information on weights and measures the different states of matter atomic and molecular weights and electron charge Giving support to these discussions are experiments that show the changes in weight and electron charge of metals gases and other materials when exposed to different conditions The text also looks at experiments on the gravimetric and volumetric stoichiometry of chlorides sulfates acids antimony and oxalates The manual also highlights studies conducted on potassium nitrate and chlorate oxygen hydrogen and polymers The guidebook ends with discussions on molecular geometry kinetics and chemical equilibrium Experiments and illustrations of chemical reactions are presented Taking into consideration the value of data presented the manual is a great find for readers wanting to introduce an organized system in conducting laboratory experiments

**SAT Subject Test Chemistry** Joseph A. Mascetta, Mark Kernion, 2020-12-01 Note College Board has discontinued the SAT Subject Tests in the US The tests will be available outside the US in June 2021 and

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### **OECD Guidelines for the Testing of Chemicals**

**OECD Guidelines for the Testing of Chemicals Fifteenth Addendum - 2004** OECD, 2004-10-07 The OECD Guidelines for the Testing of Chemicals are a collection of internationally agreed methods for testing the safety of chemicals and chemical preparations including pesticides and industrial chemicals. This 2005 update contains 11 new and 3 revised guidelines.

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**Cracking the SAT II** Princeton Review (Firm), 2003-03-11 Our inventive strategies and unique approach to standardized tests have revolutionized the test prep industry and made our SAT course the biggest in the world in just a few short years. The techniques we teach in our courses are in this book. And if it's on the SAT II French Subject Test, it's in this book. The Princeton Review realizes that acing the SAT II French Subject Test is very different from getting straight A's in school. We don't try to teach you everything there is to know about French; only the techniques you'll need to score higher on the exam.

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**Hazardous and Industrial Solid Waste Testing and Disposal** Douglas Lorenzen, American Society for Testing and Materials, 1986  
**An Introduction to Qualitative Chemical Analysis and the Related Chemical Principles** Donald Pritchard Smith, Henry Keppele Miller, 1928

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Laboratory Manual of Inorganic Chemistry John Bernard Ekeley, 1923  
Tests of the Fatigue Strength of Cast Steel University of Illinois (Urbana-Champaign campus). Engineering Experiment Station, Alfred Copeland Callen, Crandall Zachariah Rosecrans, Cloyd Moffett Smith, Herbert Fisher Moore, 1926  
**Laboratory Manual for General Chemistry** Alvin Strickler, 1923

**Laboratory Manual in General College Chemistry** Joseph Albert Babor, Alexander Lehrman, 1929  
**A Laboratory Manual in General Chemistry** William Foster, 1923  
**Illustrated Guide to Home Chemistry Experiments** Robert Bruce Thompson, 2012-02-17  
For students, DIY hobbyists and science buffs who can no longer get real chemistry sets, this one-of-a-kind guide explains how to set up and use a home chemistry lab with step-by-step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work. Purify alcohol by distillation. Produce hydrogen and oxygen gas by electrolysis. Smelt metallic copper from copper ore or make yourself Analyze the makeup of seawater, bone and other common substances. Synthesize oil of wintergreen from aspirin and rayon fiber from paper. Perform forensics tests for fingerprints, blood, drugs and poisons and much more. From the 1930s through the 1970s, chemistry sets were among the most popular Christmas gifts, selling in the millions. But two decades ago, real chemistry sets began to disappear as manufacturers and retailers became concerned about liability. The **Illustrated Guide to Home Chemistry Experiments** steps up to the plate with lessons on how to equip your home chemistry lab, master laboratory skills and work safely in your lab. The bulk of this book consists of 17 hands-on chapters that include multiple laboratory sessions on the following topics: Separating Mixtures, Solubility and Solutions, Colligative Properties of Solutions, Introduction to Chemical Reactions, Stoichiometry, Reduction, Oxidation, Redox Reactions, Acid-Base Chemistry, Chemical Kinetics, Chemical Equilibrium and Le Chatelier's Principle, Gas Chemistry, Thermochemistry and Calorimetry, Electrochemistry, Photochemistry, Colloids and Suspensions, Qualitative Analysis, Quantitative Analysis, Synthesis of Useful Compounds, Forensic Chemistry. With plenty of full-color illustrations and photos, **Illustrated Guide to Home Chemistry Experiments** offers introductory level sessions suitable for a middle school or first year high school chemistry laboratory course and more advanced sessions suitable for students who intend to take the College Board Advanced Placement AP

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Roeber,Howard Coon Parmelee,1922 **A Thermodynamic Analysis of Gas Engine Tests** Crandall Zachariah

Rosecrans,George Theodore Felbeck,1925 **The ETS Test Collection Catalog** Educational Testing Service. Test Collection,1993 The major source of information on the availability of standardized tests Wilson Library BulletinCovers commercially available standardized tests and hard to locate research instruments *Laboratory Manual of General Chemistry* Arthur Becket Lamb,1916

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to solve a linear programming problem we first need to know the fundamental theorem of linear programming given that an optimal solution to a linear programming problem exists it must occur at a vertex of the feasible set if the optimal solution occurs at two adjacent vertices of the feasible set then the linear programming problem

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Exercise 9: An optimal control problem with an analytical solution. We consider the problem of maximizing a linear function of the final state of a linear system subject to bounds on the inputs. Maximize  $\dot{x}(t)^\top u(t)$  subject to  $\dot{x}(t) = Ax(t) + Bu(t)$ ,  $x(0) = x_0$ ,  $u(t) \in [u_{\min}, u_{\max}]$ , and  $x(t) \in [x_{\min}, x_{\max}]$ . The cost function is  $J = \frac{1}{2} x(T)^\top Q x(T) + \int_0^T \frac{1}{2} u(t)^\top R u(t) dt$ .

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2018 2019 exercise 3 determine using the simplex algorithm with bland's rule the optimal solution to the following linear programming problem  $\max s.t. x_1 \leq 3x_2 \leq 5x_3 \leq 2x_4 \leq x_1 \leq 2x_2 \leq 3x_3 \leq x_4 \leq 3 \leq 2x_1 \leq x_2 \leq x_3 \leq 2x_4 \leq 4 \leq x_1 \leq x_2 \leq x_3 \leq x_4 \leq 2$  r solution the problem in standard form is the initial tableau is  $\min s.t. s_1 \leq s_2$

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recognize the typical form of a linear programming problem formulate maximization linear programming problems graph feasible regions for maximization linear programming problems determine optimal solutions for maximization linear programming problems

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nov 5 1998 as a linear program a computer program can be used to solve the problem in this regard solving a linear program is relatively easy the hardest part about applying linear programming is formulating the problem and interpreting the solution linear equations all of the equations and inequalities in a linear program must by definition be

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mathrm x 2 3 mathrm x 3 leq 18

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