

Unit 3 ~ Learning Guide Name: Aanya Nilin Shetty

Instructions:

Using a pencil, complete the following notes as you work through the related lessons. Show ALL work as is explained in the lessons. You are required to have this package completed BEFORE you write your unit test. Do your best and ask questions if you don't understand anything!

Please print out the Formula Sheet found at the start of the course. You will need this to do assignments and tests!

* Highly Exothermic reaction are less likely to be in equilibrium as they have a greater E_a reverse

Equilibrium:

1. When does equilibrium occur?

Equilibrium occurs when opposing processes occur at equal rates in a closed system at a constant temperature

* chemical equilibrium is also known as dynamic equilibrium

2. Give an example of two physical process that oppose each other during equilibrium?

A water solution in a closed container can have a physical equilibrium, and the opposing processes would be evaporation and condensation.

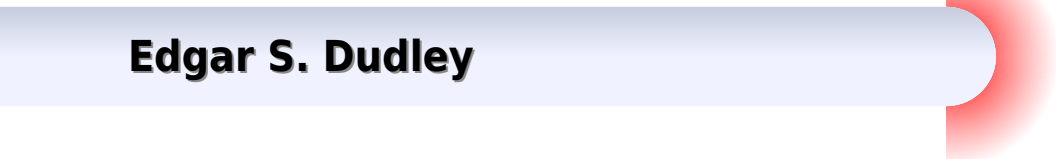
3. Complete the ICE chart for this reaction given that 4.0 moles of N_2 gas and 8.0 moles of H_2 gas are placed into a 1.00 L sealed flask and allowed to come to equilibrium. At equilibrium, 2.4 moles of NH_3 are present.

	$N_{2(g)}$	+	$3H_{2(g)}$	$\rightleftharpoons$	$2NH_{3(g)}$
Initial concentration [I]	4.0 M		8.0 M		0 M
Change in concentration [C]	-1.2 M		-3.6 M		+2.4 M
Equilibrium concentration [E]	2.8 M		4.4 M		2.4 M

* Ratio of the balanced equation (1:3:2)

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