

**Chemistry 11****Some Study Materials for the Final Exam**

Prefix	Abbreviation	Exponent
giga	G	$10^9$
mega	M	$10^6$
kilo	k	$10^3$
hecto	h	$10^2$
deca	da	$10^1$
deci	d	$10^{-1}$
centi	c	$10^{-2}$
milli	m	$10^{-3}$
micro	$\mu$	$10^{-6}$
nano	n	$10^{-9}$
pico	p	$10^{-12}$

**Complex Unit Conversion**

$$5.6 \mu\text{g/mL} = ? \text{ ng/L}$$

$$5.6 \times 10^6 \frac{\text{ng}}{\text{mL}} \times \frac{10^{-6} \text{ g}}{1 \text{ mg}} \times \frac{1 \text{ ng}}{10^{-9} \text{ g}} \times \frac{1 \text{ mL}}{10^{-3} \text{ L}} = 5.6 \times 10^6 \frac{\text{ng}}{\text{L}}$$

**Density****Precision**

- The number of significant digits to which a value has been reliably measured.
- The more decimal places, the more precise the measurement.
- The greater the precision, the less the uncertainty.

The **accuracy** of a measurement is how close the measurement is to an *accepted standard*.

**Significant Digits** Rules:

1. All non-zero digits are significant
2. Zero's between non-zero digits are significant
3. All zero's to the LEFT of the 1st non-zero digit are NOT significant. (called leading zero's)
4. Zero's to the RIGHT of the last non-zero digit ARE significant IF the decimal point is SHOWN.
5. Zero's to the RIGHT of the last non-zero digit (trailing zero's) ARE NOT significant if there is an UNDERSTOOD decimal point.
6. In Scientific Notation, the exponent part of the number is NOT significant.

# Chemistry 11 Final Exam Study Guide

**J. Compton, Allyn Jones**



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