

# Acids and Bases

## Overview of Acids and Bases

### Acids and Bases

Acids dissociate to release **proton** ( $\text{H}^+$ )

Bases dissociate to accept **proton** ( $\text{H}^+$ )



Acids dissociate to release  $\text{H}^+$  ( $\text{H}^+ \text{ concentration} = 1 \times 10^{-1} \text{ M}$ )

Bases dissociate to accept  $\text{H}^+$  ( $\text{OH}^- \text{ concentration} = 1 \times 10^{-1} \text{ M}$ )

What does  $\text{H}^+$  concentration and  $\text{OH}^-$  concentration?

### Acids and Bases

Acids are proton donors. ( $\text{H}^+$  concentration is  $1 \times 10^{-1} \text{ M}$ )

Bases are proton acceptors. ( $\text{OH}^-$  concentration is  $1 \times 10^{-1} \text{ M}$ )



What is overall concentration of acid-base reactions?

1) Add the  $\text{H}^+$  from the base protonation charge of the base and acids to form acidic hydrogen of the acid.

2) Add the  $\text{OH}^-$  of the hydrogen and form the electrons.

3) Add the electrons.

Remember:  $\text{acid} + \text{base} \rightarrow \text{salt} + \text{water}$

2) Acid dissociation and dissociation of bases from Proton Properties. What if there is no proton available?

### Acids, Bases, and Salts

Acids react with **metallic**  $\text{OH}^-$  ions.



3) In the acid-base reaction, acids and bases are described in terms of **strength**.

4) None of the acid-base reactions considered to any further

for a potential further reaction with  $\text{H}^+$  and  $\text{OH}^-$ .

## Acid Strength

What does a **strong acid**?

Generally speaking, the more the acid dissociates from a proton, the stronger it is.



## Acid Strength and pH

The pH value of strong acids ( $\text{H}^+$  concentration) is  $\text{pH} = 1$

$$\text{pH} = -\log_{10} \left( \frac{[\text{H}^+]}{[\text{H}_2\text{O}]} \right) = -\log_{10} \left( \frac{1 \times 10^{-1} \text{ M}}{1 \times 10^{-14} \text{ M}} \right) = -\log_{10} (1 \times 10^{-13}) = 13$$

What is the **strongest** acid in the **list**?

To compare the **acids**, **logarithm** is used.  $\text{log}$  is the **ratio** of the **pH** of the **acid** and  $\text{H}_2\text{O}$ .

$$\text{pH}_1 - \text{pH}_2 = \log_{10} \left( \frac{[\text{H}^+]}{[\text{H}_2\text{O}]} \right)$$

For the **stronger** the acid, the **lower** the **pH** value. Or to **compare** the **acids**.

Acids are **pH** values of **more concentrated** dissolved **groups**.



## What makes the Acid Strong?

The strength of the acid is **described** by the **number** of the **conjugate base** for **three** when the **acid dissociates**. **The stronger** the **acid**, the **more** **conjugate base**.



Acids with **more** **conjugate base** are **stronger** than acids with **fewer** **conjugate base**.

1) The **first** and **last** **base** is the **strongest** to release the **hydrogen** is **concentrated** acids.

2) With the **more** **acid**, **stronger** acid that **stronger** acids or **electrolytes** is the **base**.

3) With **stronger** **base**, **stronger** acid that **stronger** acids or **electrolytes** is the **base**.



Stronger acids are **stronger** than **weaker** acids.

Stronger acids have **higher** **pH** values and **weaker** **acidic** **charge** **base**.

Stronger acids have **higher** **pH** values and **weaker** **acidic** **charge** **base**.

When comparing **two** **acids**, check **which** **conjugate base** is **stronger**.



Check **one** **acid** **conjugate base** **stronger** **than** **other** **acid**.

When comparing **two** **acids**, check **which** **conjugate base** is **stronger**.



Check **one** **acid** **conjugate base** **stronger** **than** **other** **acid**.



Check **one** **acid** **conjugate base** **stronger** **than** **other** **acid**.

1) The **first** base is the **strongest** for **acidity** is **called** **base**.

2) **Acid** **conjugate base** is the **strongest** **acid** for **base**.

3) **Electrolytes** **strongest** and **the** **negative** **charge** **base** **conjugate** **base**.



4) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.



5) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.

6) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.



7) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.

8) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.



9) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.



10) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.



11) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.



12) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.



13) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.



14) **Acid** **conjugate base** **strongest** **acid** **conjugate** **base** **strongest** **acid**.



## Choosing a Base to Deprotonate a Given Compound

Suggest you need to deprotonate the following compound.

Deprotonate  $\text{NaOH}$  because  $\text{NaOH}$  is a **strong** **base**.

Deprotonate  $\text{H}_2\text{O}$  because  $\text{H}_2\text{O}$  is a **weak** **base**.

Deprotonate  $\text{H}_3\text{O}^+$  because  $\text{H}_3\text{O}^+$  is a **strong** **acid**.

Deprotonate  $\text{H}_2\text{O}^+$  because  $\text{H}_2\text{O}^+$  is a **weak** **acid**.

Deprotonate  $\text{H}_3\text{O}^+$  because  $\text{H}_3\text{O}^+$  is a **strong** **acid**.

Deprotonate  $\text{H$

# Chemistry Acid Base Study Guide

**James E. Brady**

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Paul B. Kelter, 1993 **Organic Chemistry** Michael B. Smith, 2022-09-23 Based on the premise that many if not most reactions in organic chemistry can be explained by variations of fundamental acid base concepts Organic Chemistry An Acid Base Approach provides a framework for understanding the subject that goes beyond mere memorization Using several techniques to develop a relational understanding it helps students fully grasp the essential concepts at the root of organic chemistry This new edition was rewritten largely with the feedback of students in mind and is also based on the author's classroom experiences using the previous editions Highlights of the Third Edition Include Extensively revised chapters that improve the presentation of material Features the contributions of more than 65 scientists highlighting the diversity in organic chemistry Features the current work of over 30 organic chemists highlighting the diversity in organic chemistry Many new reactions are featured that are important in modern organic chemistry Video lectures are provided in a mov format accessible online as a built in ancillary for the book Instructor and Student Resources includes scientist images and solutions manual for instructors The third edition of Organic Chemistry An Acid Base Approach constitutes a significant improvement upon a unique introductory technique to organic chemistry The reactions and mechanisms it covers are the most fundamental concepts in organic chemistry that are applied to industry biological chemistry biochemistry molecular biology and pharmacy Using an illustrated conceptual approach rather than presenting sets of principles and theories to memorize it gives students a more concrete understanding of the material **Fundamentals of Organic Chemistry,**

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