

Study Guide for Chapter 7 – Chemical Quantities

Test Setup: Multiple Choice, Matching, Problems: Find molar mass, determine % composition, Molar conversions will include all of the following: moles-particles, moles-mass, and moles-volume, determine the empirical formulas & molecular formulas by looking at actual formulas.

Sections in the Book:

7.1 – The Mole: A Measurement of Matter

Objectives:

- Describe how Avogadro's number is related to a mole of any substance
- Calculate the mass of a mole of any substance

Vocabulary

- mole (mol), Avogadro's number, representative particle

Practice Problems:

1. Find the molar mass of Li_2S .
2. Find the molar mass of $\text{Ca}(\text{OH})_2$.
3. How many moles in 1.50×10^{23} molecules NH_3 .
4. How many formula units in 6.78 moles of sodium chloride (NaCl)

7.2 – Mole-Mass and Mole-Volume Relationships

Objectives:

- Use the molar mass to convert between mass and moles of a substance.
- Use the mole to convert among measurements of mass and volume of substances.

Vocabulary:

- molar mass, standard temperature and pressure (STP), molar volume

Practice Problems:

1. What is the mass in grams of 0.160 mol H_2O_2 ?
2. How many moles are in 508.4 g $\text{Ca}(\text{NO}_3)_2$?
3. Calculate the volume, in liters, of 8.75 mol of chlorine gas at STP.
4. How many moles are in a container filled with 45.9 L of oxygen gas at STP?

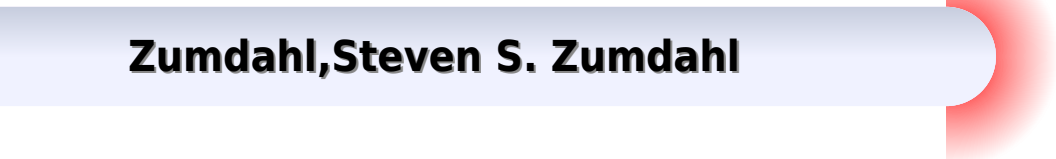
7.3 – Percent Composition and Chemical Formulas

Objectives:

- Calculate the percent composition of a substance from its chemical formula or experimental data.
- Derive the empirical formula and the molecular formula of a compound from experimental data.

Chemistry Chemical Quantities Study Guide

Zumdahl, Steven S. Zumdahl



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Study Guide for Introductory Chemistry Iris Stovall, 1990 Study Guide for Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry Iris Stovall, Steven S. Zumdahl, 1996 **Study guide to accompany chemistry** Karen Timberlakes, 1983 *Study Guide and Glossary to Accompany Chemical Principles* Daniel L. Reger, Edward E. Mercer, Robert S. Boikess, 1978 *Fundamentals of Chemistry, Study Guide* James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs **Study Guide to Accompany Basics for Chemistry** Martha Mackin, 1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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