

# AP Chemistry Midterm Study Packet

## Key Terms:

### Chapter 14 – Liquids and Solids

- amorphous solid** — solid in which the arrangement of the representative particles lacks a regular, repeating pattern; do not have a definite melting point
- boiling point** — temperature at which the vapor pressure of a liquid becomes equal to atmospheric pressure
- condensation** — change of state from liquid to gas
- condensed state** — liquid and solid states of matter; states in which a substance has a substantially higher density than in the gaseous state
- covalent network solid** — solid in which strong covalent bonds form a network extending throughout the solid
- crystalline solid** — solid in which the representative particles are in a highly ordered, repeating pattern called a crystal
- deposition** — conversion of a gas directly into a solid without first becoming a liquid
- dipole-dipole force** — intermolecular force of attraction between neighboring permanent dipoles that is created when the dipoles line up so that the positive and negative ends are close to each other
- dispersion force** — intermolecular force of attraction between induced dipoles
- equilibrium vapor pressure** — pressure exerted by a vapor in equilibrium with its liquid; point at which the number of molecules in the vapor state remains solid
- evaporation** — process by which molecules of a liquid escape from the surface of the liquid and enter the gaseous, or vapor, state
- freezing point** — temperature at which the solid and liquid form of a substance exist in equilibrium
- heat of fusion** — heat necessary to convert a given amount of a solid into a liquid
- heat of vaporization** — amount of heat necessary to vaporize a given amount of liquid
- heating curve** — plot of the temperature of a sample as a function of time
- hydrogen bond** — strong intermolecular force between the hydrogen atom of one molecule and a highly electronegative atom of another molecule
- induced dipole** — dipole created by the presence of a neighboring dipole
- intermolecular force** — relatively weak force of attraction that exists between neighboring molecules
- intramolecular force** — force of attraction that exists within a molecule to hold it together
- phase diagram** — plot of the state of a sample as a function of temperature and pressure
- sublimation** — conversion of a solid directly into a gas, without first becoming a liquid
- surface tension** — imbalance of attractive forces at the surface of a liquid that causes the surface to behave as if it had a thin film across it; resistance of a liquid to an increase in the surface area
- vaporization** — change of state from liquid to gas
- viscosity** — resistance to motion that exists between the molecules of a liquid when they move past each other; affected by the extent of the hydrogen bonding

### Chapter 15 – Solutions

- alloy** — solid solution in which the atoms of two or more metals are uniformly mixed
- aqueous solution** — solution in which the solvent is water
- boiling point elevation** — colligative property in which the boiling point of a solvent is raised when a nonvolatile solute is dissolved in the solvent; directly related to the concentration of the solution
- colligative property** — property of a solution that depends upon the concentration of the solute, but not its identity
- concentration** — amount of solute dissolved in a given amount of solvent
- freezing point depression** — colligative property in which the freezing point of a solvent is lowered when a nonvolatile solute is dissolved in the solvent; freezing point depression is directly related to the concentration of a solution
- hydration** — process in which water molecules pull solute particles into solution and form a sphere around them
- immiscible** — inability of a liquid to form a solution with another liquid in all proportions
- insoluble** — inability of a substance to dissolve in another substance
- miscible** — ability of a liquid to form a solution with another liquid in all proportions
- molarity** — concentration of a solution determined by the number of moles of solute per kilogram of solvent
- molarity** — concentration of a solution determined by the number of moles of solute per liter of solution
- mole fraction** — concentration of a solution determined by the number of moles of a given solute divided by the total number of moles of solution
- osmosis** — flow of solvent particles from a dilute solution to a concentrated solution
- osmotic pressure** — pressure required to prevent osmosis; arises when the flow of solvent particles from a dilute solution to a concentrated solution across a semipermeable membrane results in uneven heights of the solutions of either side of the membrane
- saturated** — description of a solution that contains as much dissolved solute as it can under existing conditions
- solubility** — amount of solute that dissolves in a given amount of solvent at a given temperature to form a saturated solution
- soluble** — ability of a solution to dissolve in another substance
- solute** — substance that is dissolved in a solvent to form a solution
- solution** — homogeneous mixture of two or more substances in a single physical state
- solvation** — process in which solvent particles pull solute particles into solution and form a sphere around them; called hydration if the solvent is water
- solvent** — substance that does the dissolving in a solution

# Chemistry Midterm Study Guide

**E. C. Blossey**



## Chemistry Midterm Study Guide:

*Student Study Guide* Peter L. Kresan, Reed Mencke, Frank Press, Raymond Siever, 2003-09-25 This reconceptualization of the text *Understanding Earth* reflects the fundamental changes in the field of physical geology over the past several years

*PSI Study Guide for Organic Chemistry* by Douglas C. Neckers, Michael P. Doyle E. C. Blossey, 1977 **McGraw-Hill's 500 College Chemistry Questions** David E. Goldberg, 2012-09-07 500 Ways to Achieve Your Best Grades We want you to succeed on your college chemistry midterm and final exams That's why we've selected these 500 questions to help you study more effectively use your preparation time wisely and get your best grades These questions are similar to the ones you'll find on a typical college exam so you will know what to expect on test day Each question includes comprehensive explanations in the answer key Whether you have been studying all year or are doing a last minute review McGraw Hill's 500 College Chemistry Questions will help you achieve the final grade you desire Sharpen your subject knowledge and build your test taking confidence with 500 essential college chemistry questions with answers Clear solutions in the answer key for every problem Coverage from atomic mass to electrochemistry *Teaching and Learning STEM* Richard M. Felder, Rebecca Brent, 2016-02-22 Rethink traditional teaching methods to improve student learning and retention in STEM Educational research has repeatedly shown that compared to traditional teacher centered instruction certain learner centered methods lead to improved learning outcomes greater development of critical high level skills and increased retention in science technology engineering and mathematics STEM disciplines *Teaching and Learning STEM* presents a trove of practical research based strategies for designing and teaching STEM courses at the university community college and high school levels The book draws on the authors extensive backgrounds and decades of experience in STEM education and faculty development Its engaging and well illustrated descriptions will equip you to implement the strategies in your courses and to deal effectively with problems including student resistance that might occur in the implementation The book will help you Plan and conduct class sessions in which students are actively engaged no matter how large the class is Make good use of technology in face to face online and hybrid courses and flipped classrooms Assess how well students are acquiring the knowledge skills and conceptual understanding the course is designed to teach Help students develop expert problem solving skills and skills in communication creative thinking critical thinking high performance teamwork and self directed learning Meet the learning needs of STEM students with a broad diversity of attributes and backgrounds The strategies presented in *Teaching and Learning STEM* don't require revolutionary time intensive changes in your teaching but rather a gradual integration of traditional and new methods The result will be continual improvement in your teaching and your students learning More information about *Teaching and Learning STEM* can be found at <http://educationdesignsinc.com> book including its preface foreword table of contents first chapter a reading guide and reviews in 10 prominent STEM education journals *Making Pre-Med Count* Elisabeth Fassas, 2020-01-07 In *Making Pre Med Count* med student Elisabeth Fassas shares

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### **Table of Contents Chemistry Midterm Study Guide**

1. Understanding the eBook Chemistry Midterm Study Guide
  - The Rise of Digital Reading Chemistry Midterm Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Midterm Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Midterm Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Midterm Study Guide
  - Personalized Recommendations
  - Chemistry Midterm Study Guide User Reviews and Ratings
  - Chemistry Midterm Study Guide and Bestseller Lists

5. Accessing Chemistry Midterm Study Guide Free and Paid eBooks
  - Chemistry Midterm Study Guide Public Domain eBooks
  - Chemistry Midterm Study Guide eBook Subscription Services
  - Chemistry Midterm Study Guide Budget-Friendly Options
6. Navigating Chemistry Midterm Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Midterm Study Guide Compatibility with Devices
  - Chemistry Midterm Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Midterm Study Guide
  - Highlighting and Note-Taking Chemistry Midterm Study Guide
  - Interactive Elements Chemistry Midterm Study Guide
8. Staying Engaged with Chemistry Midterm Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Midterm Study Guide
9. Balancing eBooks and Physical Books Chemistry Midterm Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Midterm Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Midterm Study Guide
  - Setting Reading Goals Chemistry Midterm Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Midterm Study Guide
  - Fact-Checking eBook Content of Chemistry Midterm Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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