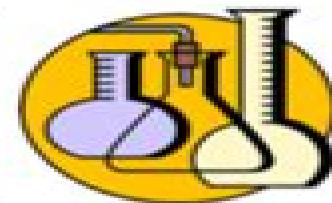


**Stoichiometry and The Mole – Review Guide****Remember:**

- 1 mole =  $6.02 \times 10^{23}$  atoms, molecules, formula units
- 1 mole = grams (of molar mass)
- 1 mole = 22.4 liters (L) of gas



- 1) The process used to predict the amount of a product that will be formed is **stoichiometry**
- 2) If you know the mass of a substance, you can find the number of moles by dividing it by the...  
**molar mass of substance**
- 3) If you know the number of particles (atoms) of a substance, you can find the number of moles by dividing it by...

**Avogadro's Number**

- 4) Summarize the steps of a gram-to-gram stoichiometry problem:

Grams A → **Moles A** → **Moles B** → Grams B

- 5) The essential stoichiometry problem needed to convert from one chemical to another is:

**Mole** to **Mole**

- 6) Two reactants undergo a chemical reaction. At the end of the reaction, Chemical A still exists in the container, while Chemical B has been completely used up. Which is the limiting reactant?

**Chemical B**

- 7) What is a mole? **A unit of measurement used to count atoms/molecules.**

- 8) What is Avogadro's Number?  **$6.02 \times 10^{23}$**

- 9) What does Avogadro's number tell us? **The number of atoms or molecules in a mole.**

# Chemistry Moles Study Guide

**Allan K. Fleeger, Dean R. Lillquist**

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