

Name _____ Date _____ Class _____

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

Solutions

Section 15.1 What are solutions?

In your notebook, read about the characteristics of solutions.

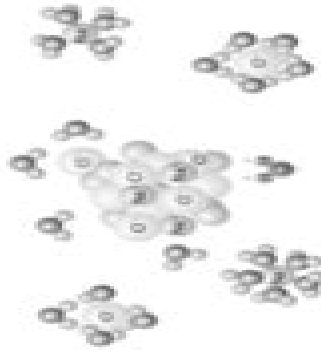
Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
miscible	insoluble	solvent	solvent

Air is an (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen in air is the (2) solute and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is an (4) soluble substance. A substance that does not dissolve is (5) insoluble. (6) Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



9. Hydration is solvation in which the solvent is water. What is solvation?
Solvation is the process of surrounding solute particles with solvent particles to form a solution.

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Chemistry: Matter and Change • Chapter 15

85

Name _____ Date _____ Class _____

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?
The sodium and chloride ions are separated and surrounded by the water molecules.
11. Explain the orientation of the water molecules around the sodium ions and chloride ions.
Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.
12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?
The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.
13. List three ways that the rate of solvation may be increased.
stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your notebook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write *true* or *false*.

- true 14. The overall energy change that occurs when a solution forms is called the heat of solution.
- false 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.
- false 16. Solvation continues as long as the solvation rate is less than the crystallization rate.
- true 17. In a saturated solution, solvation and crystallization are in equilibrium.
- true 18. Additional solute can be dissolved in an unsaturated solution.
- true 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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86

Chemistry Solutions Study Guide Answer Key

Kenneth W. Raymond



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Study Guide and Solutions Manual to accompany Basic Concepts of Chemistry, 9th Edition Leo J. Malone, 2011-12-07 This is the Study Guide and Solutions Manual to accompany Malone's Basic Concepts of Chemistry

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 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Solutions Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Solutions Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Solutions Study Guide Answer Key
 - Personalized Recommendations
 - Chemistry Solutions Study Guide Answer Key User Reviews and Ratings
 - Chemistry Solutions Study Guide Answer Key and Bestseller Lists
5. Accessing Chemistry Solutions Study Guide Answer Key Free and Paid eBooks
 - Chemistry Solutions Study Guide Answer Key Public Domain eBooks
 - Chemistry Solutions Study Guide Answer Key eBook Subscription Services
 - Chemistry Solutions Study Guide Answer Key Budget-Friendly Options

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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Solutions Study Guide Answer Key
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Solutions Study Guide Answer Key
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Chemistry Solutions Study Guide Answer Key
 - Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of Chemistry Solutions Study Guide Answer Key
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 - Utilizing eBooks for Skill Development
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