

**CHAPTER 15** STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

is your method, read about the characteristics of solutions. Use each of the terms below just once to complete the sentence.

solvent	aqueous	solute	solubility
soluble	soluble	soluble	soluble

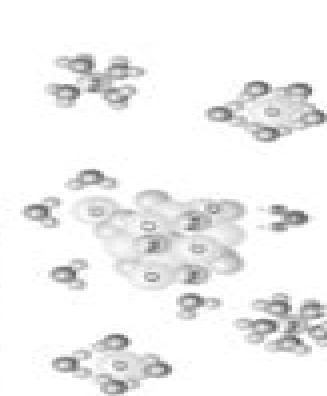
After it is in (1) \_\_\_\_\_, a solute \_\_\_\_\_ is dissolved in a solvent \_\_\_\_\_.

(2) \_\_\_\_\_ is a solution in which the (3) \_\_\_\_\_ is the (4) \_\_\_\_\_, and the solute is the (5) \_\_\_\_\_. In some (6) \_\_\_\_\_, the solute is dissolved in a solvent, and the (7) \_\_\_\_\_ is the solvent. A solution that does not dissolve is (8) \_\_\_\_\_.

If one liquid is soluble in another liquid, such as water and oil, water is (9) \_\_\_\_\_ in (10) \_\_\_\_\_. However, if one liquid is insoluble in another, the liquids are (11) \_\_\_\_\_.

Read about solubility in aqueous solutions in your textbook.

Use the diagram to answer the following questions.



The diagram shows the hydration of water molecules that have an aqueous solution.

Use the diagram to answer the following questions.

1. Indicate in solution in which the solute is water. What is a solute?

Solution is the process of surrounding solute particles with solvent particles to form a solution.

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**STUDY GUIDE FOR CONTENT MASTERY****Section 15.1** *continued*

16. In sodium chloride dissolved in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are **separated** and **surrounded by the water molecules**.

17. Explain the attraction of the water molecules around the sodium ion and chloride ion.

Because the sodium ion is **positively charged**, it **attracts the negatively charged portion of the water molecule (the oxygen atom)** and **repels the positively charged portion of the water molecule (the hydrogen atoms)**, because the chloride ion is **negatively charged**, it **attracts the positively charged portion of the water molecule** and **repels the negatively charged portion**.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How is it possible?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a pure aqueous solution, the concentration of solute may be increased by stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent.

In pure (20) \_\_\_\_\_, read about heat of solution, miscibility, and (21) \_\_\_\_\_ that affect solubility for each (22) \_\_\_\_\_ below, write true or false.

True \_\_\_\_\_ 16. The overall ionic charge that occurs when a solute dissolves in a solvent is called the heat of solution.

False \_\_\_\_\_ 17. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

False \_\_\_\_\_ 18. Solvent molecules in a pure solution are in constant motion.

True \_\_\_\_\_ 19. In a saturated solution, solute and solvent molecules are in equilibrium.

True \_\_\_\_\_ 20. Additional solute can be dissolved in an unsaturated solution.

True \_\_\_\_\_ 21. The solubility of gas dissolved in a liquid decreases as the temperature of the solution increases.

True \_\_\_\_\_ 22. Dissolving Water and Oils *• Chapter 15*

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# Chemistry Solutions Study Guide Answer Key

**Kenneth W. Raymond**

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