

STAAR CHEMISTRY REFERENCE MATERIALS



ATOMIC STRUCTURE

$$\text{Speed of light} = (\text{frequency})(\text{wavelength})$$

$$c = f\lambda$$

$$\text{Energy} = (\text{Planck's constant})(\text{frequency})$$

$$E_{\text{photon}} = hf$$

$$\text{Energy} = \frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})}$$

$$E_{\text{photon}} = \frac{hc}{\lambda}$$

BEHAVIOR OF GASES

$$\text{Total pressure of a gas} = \left(\frac{\text{sum of the partial pressures}}{\text{of the component gases}} \right)$$

$$P_T = P_1 + P_2 + P_3 + \dots$$

$$(\text{Pressure})(\text{volume}) = (\text{moles})(\text{ideal gas constant})(\text{temperature})$$

$$PV = nRT$$

$$\frac{(\text{initial pressure})(\text{initial volume})}{(\text{initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})}$$

$$\frac{PV_1}{n_1T_1} = \frac{PV_2}{n_2T_2}$$

$$(\text{initial pressure})(\text{initial volume}) = (\text{final pressure})(\text{final volume})$$

$$P_1V_1 = P_2V_2$$

$$\frac{(\text{initial volume})}{(\text{initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})}$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{(\text{initial volume})}{(\text{initial moles})} = \frac{(\text{final volume})}{(\text{final moles})}$$

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

SOLUTIONS

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liter of solution}}$$

$$M = \frac{\text{mol}}{\text{L}}$$

$$\text{Ionization constant of water} = \left(\frac{\text{hydrogen ion}}{\text{concentration}} \right) \left(\frac{\text{hydroxide ion}}{\text{concentration}} \right)$$

$$K_w = [\text{H}^+][\text{OH}^-]$$

$$\left(\frac{\text{volume of}}{\text{solution 1}} \right) \left(\frac{\text{molarity of}}{\text{solution 1}} \right) = \left(\frac{\text{volume of}}{\text{solution 2}} \right) \left(\frac{\text{molarity of}}{\text{solution 2}} \right)$$

$$V_1M_1 = V_2M_2$$

$$\text{pH} = \text{-logarithm (hydrogen ion concentration)}$$

$$\text{pH} = -\log[\text{H}^+]$$

THERMOCHEMISTRY

$$\text{Heat gained or lost} = (\text{mass}) \left(\frac{\text{specific}}{\text{heat}} \right) \left(\frac{\text{change in}}{\text{temperature}} \right)$$

$$Q = mc_p \Delta T$$

$$\text{Enthalpy of reaction} = \left(\frac{\text{enthalpy}}{\text{of products}} \right) - \left(\frac{\text{enthalpy}}{\text{of reactants}} \right)$$

$$\Delta H = \Delta H_f^\circ(\text{products}) - \Delta H_f^\circ(\text{reactants})$$

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