

## Colligative Properties – Chemistry 101 Study Guide

### What Are Colligative Properties?

Colligative properties are physical properties of solutions that depend only on the number of solute particles in a given quantity of solvent—not the type of particles. These are key to understanding behavior in dilute solutions.

#### 1. Vapor Pressure Lowering

When a non-volatile solute is added to a solvent, the vapor pressure of the solvent decreases. Fewer solvent molecules are at the surface, so fewer escape into the vapor phase.

Raoult's Law:  $P_{\text{solution}} = X_{\text{solvent}} \cdot P_{\text{pure solvent}}$

#### 2. Boiling Point Elevation

The boiling point of a solution is higher than that of the pure solvent. This occurs because more heat is needed to make the vapor pressure equal atmospheric pressure.

$$\Delta T_b = i \cdot K_b \cdot m$$

Where:

$\Delta T_b$  = boiling point elevation

$K_b$  = ebullioscopic constant

$i$  = van't Hoff factor

$m$  = molality

#### 3. Freezing Point Depression

The freezing point of a solution is lower than that of the pure solvent. Solute particles interfere with the formation of a solid structure.

$$\Delta T_f = i \cdot K_f \cdot m$$

Where:

$\Delta T_f$  = freezing point depression

$K_f$  = cryoscopic constant

#### 4. Osmotic Pressure

Osmotic pressure is the pressure required to stop osmosis—the movement of solvent molecules through a semipermeable membrane into a more concentrated solution.

$$\Pi = i \cdot M \cdot R \cdot T$$

Where:

$\Pi$  = osmotic pressure

$M$  = molarity

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**Thomas Elliott Taylor**

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**Clinical Chemistry Study Guide** Norbert W. Tietz, 1987

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