

# CHEMISTRY



## Periodic Table of Elements

- Elemental Forms:**
  - Monatomics:** Noble gases; helium, neon, argon, krypton, xenon, and radon; nonmetals.
  - Diatomics:** Generic fluorine, chlorine, hydrogen, oxygen, and nitrogen; Liquid bromine; Sodium iodide.
  - Molecular solids:** Sulfur
  - Covalent solids:** Carbon (diamond, graphite)
  - Metals:** Most common type of element; high coordination numbers; very dense.
- Types of covalence:** Large variety due to ionic, covalent, and metallic bonding options.
  - Basic:** Electron exchange
  - Covalent:** Shared electrons
    - Polar covalent:** Unequal sharing
    - Ion covalent:** Central ion + Solvent
  - Transition metal covalence:**
    - Coordination covalence:** Non-bonding ligands
    - Organometallics:** Metal-C bond
  - Complexes:** molecules; Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
  - Family of elements:** Columns in periodic table
  - The location of orbitals produces an expanded effect

### Periodic Trends

- Chemistry determined by the behavior of valence electrons
- Atoms or ions get larger as you go down the column; electron shells are added
- Cations and anions (same charge) decrease in size as the atomic number (Z) increases

### Practical Application

- The periodic table is used to predict the properties of elements in the periodic table.

## Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom,  $Z = \text{number of electrons}$ ; for an anion of charge  $-n$ ,  $Z = n$  electrons; for a cation of charge  $+n$ ,  $Z = n$  electrons
- Mass number (A):**  $A = Z + N$ , where  $N$  = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same  $Z$ , different  $A$ ; an element can have one or more isotopes; any sample contains a mixture of isotopes (in practical terms, use the average mass number or atomic weight (weighted average of natural isotopes for a given element))



## Atomic Quantum Numbers & Orbitals

**Pauli principle:** An angular momentum (orbital shape) and magnetic (orbital direction)  $\Rightarrow$  electron spin

- For each  $n$ , the possible  $l$  values are  $0, 1, \dots, n-1$
- For each  $l$ , the possible  $m_l$  values are  $-l, \dots, l$
- For  $m_l$ , there are possible values:  $+1/2$  and  $-1/2$  (spin up, spin down)
- Each value of  $n$  electrons is a "shell" in the atomic structure; each  $l$  electrons is a subshell
- $n$ -orbital:  $l = 0, 1$  type
- $l$ -orbital:  $l = 2, 3$  types
- $g$ -orbital:  $l = 0, 3$  types
- $f$ -orbital:  $l = 3, 7$  types

### Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe many-electron atoms and ionic; a more rigorous treatment determines exact energy levels and orbital properties that result in that subshells have different energies; the filling of levels is guided by the Aufbau principle:



| $n$ | $l$ | $m_l$ | $\ell$ | Subshell |
|-----|-----|-------|--------|----------|
| 1   | 0   | 0     | -      | 1s       |
| 2   | 0   | 0     | -      | 2s       |
| 2   | 0   | 0     | -      | 2p       |
| 2   | 1   | -1    | -      | 2p       |
| 2   | 1   | 0     | -      | 2p       |
| 2   | 1   | 1     | -      | 2p       |
| 3   | 0   | 0     | -      | 3s       |
| 3   | 1   | -1    | -      | 3p       |
| 3   | 1   | 0     | -      | 3p       |
| 3   | 1   | 1     | -      | 3p       |
| 4   | 0   | 0     | -      | 4s       |
| 4   | 1   | -1    | -      | 4p       |
| 4   | 1   | 0     | -      | 4p       |
| 4   | 1   | 1     | -      | 4p       |
| 4   | 2   | -2    | -      | 4d       |
| 4   | 2   | -1    | -      | 4d       |
| 4   | 2   | 0     | -      | 4d       |
| 4   | 2   | 1     | -      | 4d       |
| 4   | 2   | 2     | -      | 4d       |

- Electron configuration:** Atomic orbital occupancy for an atom
- Fundamental principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons (one with spin up,  $m_l = +1/2$ , and one with spin down,  $m_l = -1/2$ )
- Maxwell's rule:** Electrons fill  $p$ ,  $d$ , and  $f$  subshells in a manner maximizing the number of unpaired electrons (half-filled orbitals); spin is random!
- Ionization potential (IP):** Energy required to remove electrons from an atom or ion
  - First IP: Removal of first valence electron
  - Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale; 0 for rare gas

# College Chemistry Study Guide

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