

How to Name Chemical Compounds

is it ionic? is an anion,
or does it contain a polyatomic ion (as provided by H⁺)?

Ionic compound

[cation's name] [anion's name]

| | | |
|------------------|--|--|
| main- element | <ul style="list-style-type: none"> take its element name! ex. Al^{3+}: aluminum if transition metal*, include charge in Roman numeral ex. Cr^{2+}: chromium (II) | <ul style="list-style-type: none"> get element's name, chop end, add "-ide" ex. F^{-}: fluorine → fluoride ✓ |
| poly- element | <ul style="list-style-type: none"> just use its name! ex. NH_4^{+}: ammonium | <ul style="list-style-type: none"> just use its name! ex. CO_3^{2-}: carbonate |

* exceptions (no Roman numerals):

silver ion: always Ag^{+}

zinc ion: always Zn^{2+}

cadmium ion: always Cd^{2+}

Molecular compound

is it HCO_3^{-} ? otherwise ...

Acid! name and formula

[#]-[1st element name] [#]-[2nd element root]-"ide"

↑ skip if "mono-" ↑ skip if 1st element is H

Numerical prefixes

- | | |
|-------------|-------------|
| • 1: mono- | • 6: hexa- |
| • 2: di- | • 7: hepta- |
| • 3: tri- | • 8: octa- |
| • 4: tetra- | • 9: nona- |
| • 5: penta- | • 10: deca- |

for
first element: never
ex. N_2O_5 → dinitrogen
pentoxide

Acids

Ignore all H^{+} , and look at the name of the remaining anion.

| Anion | X-ide | X-ate | X-ite |
|-------|--|---|--|
| Acid | hydro-X-ic acid | X-ic acid | X-ous acid |
| | ex. Br^{-} bromide HBr hydrobromic acid | ex. $\text{C}_2\text{H}_3\text{O}_2^{-}$ acetate $\text{HC}_2\text{H}_3\text{O}_2$ acetic acid | ex. IO_3^{-} iodate HIO_3 iodic acid |
| | ex. CN^{-} cyanide HCN hydrocyanic acid | ex. PO_4^{3-} phosphate H_3PO_4 phosphoric acid | ex. SO_3^{2-} sulfite H_2SO_3 sulfurous acid |

Chemistry Study Guide How Elements Form Compound

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Table of Contents Chemistry Study Guide How Elements Form Compound

1. Understanding the eBook Chemistry Study Guide How Elements Form Compound
 - The Rise of Digital Reading Chemistry Study Guide How Elements Form Compound
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide How Elements Form Compound
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide How Elements Form Compound
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide How Elements Form Compound
 - Personalized Recommendations
 - Chemistry Study Guide How Elements Form Compound User Reviews and Ratings
 - Chemistry Study Guide How Elements Form Compound and Bestseller Lists
5. Accessing Chemistry Study Guide How Elements Form Compound Free and Paid eBooks

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 - Chemistry Study Guide How Elements Form Compound eBook Subscription Services
 - Chemistry Study Guide How Elements Form Compound Budget-Friendly Options
6. Navigating Chemistry Study Guide How Elements Form Compound eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Study Guide How Elements Form Compound Compatibility with Devices
 - Chemistry Study Guide How Elements Form Compound Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Study Guide How Elements Form Compound
 - Highlighting and Note-Taking Chemistry Study Guide How Elements Form Compound
 - Interactive Elements Chemistry Study Guide How Elements Form Compound
 8. Staying Engaged with Chemistry Study Guide How Elements Form Compound
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide How Elements Form Compound
 9. Balancing eBooks and Physical Books Chemistry Study Guide How Elements Form Compound
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide How Elements Form Compound
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Study Guide How Elements Form Compound
 - Setting Reading Goals Chemistry Study Guide How Elements Form Compound
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Study Guide How Elements Form Compound
 - Fact-Checking eBook Content of Chemistry Study Guide How Elements Form Compound
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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