

1

SOME BASIC CONCEPTS OF CHEMISTRY

Weightage 8 marks

1 mark - 1Q (Q No. 1)

2 mark - 1Q (Q No. 11)

5 mark - 1Q (Q No. 27)

ONE MARK QUESTIONS

1. Following medicines are used in the treatment of which disease?

- (a) Cisplatin (PUB)
 (b) Taxol (EQ)
 (c) Drug AZT (Azidothymidine) (EQ)

Ans: (a) Treatment of Cancer
 (b) Treatment of Cancer
 (c) Treatment of AIDS patients.

2. Give the S.I. unit of

- (i) Density (PUB, May-17, Feb -17,19, 20)
 (ii) Temperature (May 15)
 (iii) Amount of substance (Feb 18)
 (iv) Electric current (Feb 19)
 (v) Luminous intensity
 (vi) Volume (Feb 20)

Ans: (i) kg/m^3 (ii) Kelvin (iii) Mole
 (iv) Ampere (v) Candela (vi) m^3

3. Express the following in scientific notation

- (i) 0.00016 (Feb 18)

Ans: 1.6×10^{-4}

- (ii) 0.00085 (May 15)

Ans: 8.5×10^{-4}

- (iii) 0.00053 (Feb 16)

Ans: 5.3×10^{-4}

- (iv) 8008 (EQ)

Ans: 8.008×10^3

- (v) 0.001023 (PUB)

Ans: 1.023×10^{-3}

- (vi) 232.508 (Feb 19)

Ans: 2.32508×10^2

4. How many significant figures are there in

- (i) 0.00425 (Feb 17)
 (ii) 20.0
 (iii) 6.022×10^{23}
 (iv) 0.2500 (PUB)
 (v) 3.10087
 (vi) 8.2015

Ans: (i) 3 (ii) 3 (iii) 4 (iv) 4 (v) 6 (vi) 5

5. Give the relation between

- (a) Degree Fahrenheit and degree Celsius. (Feb 17)
 (b) Degree Celsius and Kelvin (EQ)

Ans: (a) $^{\circ}\text{F} = 9/5 \times (^{\circ}\text{C}) + 32$
 (b) $\text{K} = ^{\circ}\text{C} + 273$

6. State law of Conservation of Mass (EQ)

Ans: It states that "matter can neither be created nor destroyed".

7. State law of Definite Proportions (PUB)

Ans: The law states that a given compound always contains exactly the same proportion of elements by weight.

8. State law of 'Multiple Proportions'. (EQ)

Ans: When two elements can combine to form more than one compound, the masses of one element that combine with a fixed mass of the other element, are in the ratio of small whole numbers.

9. State 'Gay Lussac's Law' of gaseous volumes. (EQ)

Ans: The volume of gases taking part in a chemical reaction show simple whole number ratio to one another when all the volumes are measured at the same temperature and pressure.

10. State Avogadro's law. (May14, Feb-17,20)

Ans: It states that equal volumes of all gases contain equal number of molecules at same temperature and pressure.

Chemistry Study Guide Questions And Concepts

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



Chemistry Study Guide Questions And Concepts:

Basic Concepts of Chemistry, Study Guide Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples **Robinson Chemistry Study Guide** Robinson, 1992 **Study Guide with Selected Solutions** Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems [Chemistry - Study Guide to Accompany Moore/Davies/Collins](#) Clare T. Furse, 1978 **Chemistry** Richard Post, Chad Snyder, Clifford C. Houk, 2020-09-16 A practical complete and easy to use guide for understanding major chemistry concepts and terms Master the fundamentals of chemistry with this fast and easy guide Chemistry is a fundamental science that touches all other sciences including biology physics electronics environmental studies astronomy and more Thousands of students have successfully used the previous editions of Chemistry Concepts and Problems A Self Teaching Guide to learn chemistry either independently as a refresher or in parallel with a college chemistry course This newly revised edition includes updates and additions to improve your success in learning chemistry This book uses an interactive self teaching method including frequent questions and study problems increasing both the speed of learning and retention Monitor your progress with self tests and master chemistry quickly This revised Third Edition provides a fresh step by step approach to learning that requires no prerequisites lets you work at your own pace and reinforces what you learn ensuring lifelong mastery Master the science of basic chemistry with this innovative self paced study guide Teach yourself chemistry refresh your knowledge in preparation for medical studies or other coursework or enhance your college chemistry course Use self study features including review questions and quizzes to ensure that you re really learning the material Prepare for a career in the sciences medicine or engineering with the core content in this user friendly guide Authored by expert postsecondary educators this unique book gently leads students to deeper levels and concepts with practice critical thinking problem solving and self assessment at every stage **A Self-study Guide to the Principles of Organic Chemistry** Jiben Roy, 2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of

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