

# Intro. to Chemical Reactions

## Electrolyte Reactions

1. Write down formulas for the following compounds:

a) Barium hydroxide  $\text{Ba}(\text{OH})_2$

b) Sodium phosphate  $\text{Na}_3\text{PO}_4$

c) Magnesium phosphate  $\text{Mg}_3(\text{PO}_4)_2$

d) Calcium iodide  $\text{CaI}_2$

2. Write the formulas for the following acids:

a) HClO<sub>4</sub> Hydrogen perchlorate

b) H<sub>3</sub>PO<sub>4</sub> Hydrogen phosphate

c) H<sub>2</sub>SO<sub>4</sub> Hydrogen sulfate

d) H<sub>2</sub>CO<sub>3</sub> Hydrogen carbonate

## Electrolyte Reactions: Physical Theory Review

There is often a chemical reaction that does not go forward:



There are reasons that you can come up with when looking at this:

**Common/Shared** reagents with **shared** and **variable** **affinities** and **parent** **ions**

There is often no clear relationship between the common and the variable substances and there is a clear agreement:

## Common/Shared Reagents

1. Given the following reaction, write the equation that goes along with it:



Common/Shared reagents with **shared** **affinity** and **variable** **affinity** **substances** and **parent** **ions**

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**L Darling-Hammond**

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