

# Chemistry Unit 3

## Study Guide

### Chemical Nomenclature

#### Chemical formula

- a way of presenting information about the chemical proportions of atoms that constitute a particular chemical compound or molecule, using chemical element symbols, numbers, and sometimes also other symbols

#### Subscript

- A subscript or superscript is a character that is set slightly below or above the normal line of type, respectively. It is usually smaller than the rest of the text. Subscripts appear at or below the baseline, while superscripts are above.

#### Ion

- Ion, any atom or group of atoms that bears one or more positive or negative electrical charges.

#### Anion

- a negatively charged ion

#### Cation

- A positively charged ion

#### Polyatomic ion

- The prefix poly- means many, so a polyatomic ion is an ion that contains more than one atom.

#### Understand the use of Roman Numerals in ionic compounds

- Roman numerals are used in naming ionic compounds when the metal cation forms more than one ion.

#### List the diatomic molecules

- Hydrogen ( $H_2$ )
- Nitrogen ( $N_2$ )
- Oxygen ( $O_2$ )
- Fluorine ( $F_2$ )
- Chlorine ( $Cl_2$ )
- Iodine ( $I_2$ )
- Bromine ( $Br_2$ )

#### Be able to name and write chemical formulas for ionic, covalent, and acidic compounds

- Properly use prefixes
- Properly use Roman Numerals
- Properly use endings (-ide, -ate, -ite, -ous, -ic)

#### Determine the type of compound made from a pair of elements or polyatomic ions

#### Identify the ionic charge of transition metals in compounds

- Because most transition metals have two valence electrons, the charge of  $2+$  is a very common one for their ions.

#### Identify compounds as ionic, covalent, or acidic

- Compounds containing two elements (so called binary compounds) can either have ionic or covalent bonding. If a compound is made from a metal and a nonmetal, its bonding will be ionic. If a compound is made from two non-metals, its bonding will be covalent.

# Chemistry Unit 3 Study Guide Answer

**Richard Bailey**



### **Chemistry Unit 3 Study Guide Answer:**

Resources in Education ,1980      **Basic Concepts of Chemistry, Study Guide** Leo J. Malone,2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems      *Gage Canadian GED Study Guide Series* Kathy Austin,1998      **Chemistry** John S. Phillips,Cheryl Wistrom,2000      **The Publishers' Trade List Annual** ,1904      *Catalog of Copyright Entries. Third Series* Library of Congress. Copyright Office,1976      Journal of Chemical Education ,1928 *Paperbound Book Guide for Colleges* ,1972      Abstracts of Theses, Researches in Progress, and Bibliography of Publications ,1930      Catalog and Announcement of Courses San Diego State College,1899      **El-Hi Textbooks in Print** ,1984      **Organic Chemistry** Evelyn Halpern,1990      **McGraw-Hill Encyclopedia of Science and Technology** Daniel N. Lapedes,1977 Accompanied by Readers guide 33 p 28 cm New York McGraw Hill c1977 Call number Q121 M3 1977 Suppl 1 Accompanied by Study guide 50 p 28 cm New York McGraw Hill c1977 Call number Q121 M3 1977 Suppl 2

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