

# Chem II Final Exam Study Guide

## Lecture Notes

### Chapter 14 - Solutions

#### Mixtures:

- Dissolve vs dissociate: dissolve means to break apart in a solution, dissociate means to break the molecule up into ions in a solution
- Entropy (S) = a drive to disorder
  - Mixtures are more disordered than pure substances

#### Intermolecular forces:

- a. Definition: interaction between individual molecules
- b. To break IMFs = release energy (endothermic)
- c. To build IMFs = to absorb energy (endothermic)
- d. Characteristics of strong IMFs:
  - i. Higher boiling points, lower melting points, higher viscosity, lower VP
- e. 4 types of IMFs:
  - i. Dispersion: in everything. Dispersion is **stronger in larger molecules**
  - ii. Dipole-dipole: occurs in polar molecules, higher  $\Delta EN$  = higher d-d forces
  - iii. H bonds: hydrogen with N, O, F, occurs in polar molecules
- f. Polar dissolves polar, non polar dissolves non polar,
  - i. Less nonpolar shrubbery = more polar
- g. Increasing / decreasing volume in a solution = looser, weaker bonds
- h. Examples of solutions:
  - i. Ideal solution: breaks the same type of bonds that were there originally  $\Delta V = 0$ ,  $\Delta T = 0$ ,  $\Delta H = 0$
  - ii. Miscible non-ideal: break weaker, form stronger,  $\Delta V = (-)$  decrease,  $\Delta T = (+)$  increase,  $\Delta H = (-)$  exothermic
  - iii. Tends to be immiscible, non ideal: ex: oil + water, break stronger IMFs, form weaker,  $\Delta V = (+)$  increase  $\Delta T = (-)$  decrease,  $\Delta H = (+)$  endothermic

#### Concentration units:

- a. m = molality = mol solu / Kg solv
- b. M = molarity mol solu / L soln
- c. % m/v = g solu / mL soln
- d. ppm/ppb/pppt = 1 g solu /  $10^6$  g soln,  $10^9$  for ppb,  $10^{12}$  for ppt
- e. Mole fraction:  $X_a = \frac{n_a}{n_a + n_b}$

#### Henry's Law:

- Solubility of a gas when you're trying to dissolve it into a liquid
- $S = K_H \times P_{gas}$   $K_H$  is given, pressure in ATM

#### Raoult's Law:

- **Use everything you see vapor pressure**
- Tendency for a mixture to evaporate

# Chem Student Solutions Guide

**Paula Y. Bruice**



## Chem Student Solutions Guide:

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### **Table of Contents Chem Student Solutions Guide**

1. Understanding the eBook Chem Student Solutions Guide
  - The Rise of Digital Reading Chem Student Solutions Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chem Student Solutions Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chem Student Solutions Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chem Student Solutions Guide
  - Personalized Recommendations
  - Chem Student Solutions Guide User Reviews and Ratings

- Chem Student Solutions Guide and Bestseller Lists
- 5. Accessing Chem Student Solutions Guide Free and Paid eBooks
  - Chem Student Solutions Guide Public Domain eBooks
  - Chem Student Solutions Guide eBook Subscription Services
  - Chem Student Solutions Guide Budget-Friendly Options
- 6. Navigating Chem Student Solutions Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Chem Student Solutions Guide Compatibility with Devices
  - Chem Student Solutions Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chem Student Solutions Guide
  - Highlighting and Note-Taking Chem Student Solutions Guide
  - Interactive Elements Chem Student Solutions Guide
- 8. Staying Engaged with Chem Student Solutions Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chem Student Solutions Guide
- 9. Balancing eBooks and Physical Books Chem Student Solutions Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chem Student Solutions Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chem Student Solutions Guide
  - Setting Reading Goals Chem Student Solutions Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chem Student Solutions Guide
  - Fact-Checking eBook Content of Chem Student Solutions Guide
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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