

Unit 3 ~ Learning Guide Name: Aanya Nilin Shetty

Instructions:

Using a pencil, complete the following notes as you work through the related lessons. Show ALL work as is explained in the lessons. You are required to have this package completed BEFORE you write your unit test. Do your best and ask questions if you don't understand anything!

Please print out the Formula Sheet found at the start of the course. You will need this to do assignments and tests!

* Highly Exothermic reaction are less likely to be in equilibrium as they have a greater E_a reverse

Equilibrium:

1. When does equilibrium occur?

Equilibrium occurs when opposing processes occur at equal rates in a closed system at a constant temperature

* chemical equilibrium is also known as dynamic equilibrium

2. Give an example of two physical process that oppose each other during equilibrium?

A water solution in a closed container can have a physical equilibrium, and the opposing processes would be evaporation and condensation.

3. Complete the ICE chart for this reaction given that 4.0 moles of N_2 gas and 8.0 moles of H_2 gas are placed into a 1.00 L sealed flask and allowed to come to equilibrium. At equilibrium, 2.4 moles of NH_3 are present.

	$N_{2(g)}$	+	$3H_{2(g)}$	$\rightleftharpoons$	$2NH_{3(g)}$
Initial concentration [I]	4.0 M		8.0 M		0 M
Change in concentration [C]	-1.2 M		-3.6 M		+2.4 M
Equilibrium concentration [E]	2.8 M		4.4 M		2.4 M

* Ratio of the balanced equation (1:3:2)

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Table of Contents Chemistry 12 Guide

1. Understanding the eBook Chemistry 12 Guide
 - The Rise of Digital Reading Chemistry 12 Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 12 Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 12 Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 12 Guide
 - Personalized Recommendations

- Chemistry 12 Guide User Reviews and Ratings
- Chemistry 12 Guide and Bestseller Lists
- 5. Accessing Chemistry 12 Guide Free and Paid eBooks
 - Chemistry 12 Guide Public Domain eBooks
 - Chemistry 12 Guide eBook Subscription Services
 - Chemistry 12 Guide Budget-Friendly Options
- 6. Navigating Chemistry 12 Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry 12 Guide Compatibility with Devices
 - Chemistry 12 Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 12 Guide
 - Highlighting and Note-Taking Chemistry 12 Guide
 - Interactive Elements Chemistry 12 Guide
- 8. Staying Engaged with Chemistry 12 Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 12 Guide
- 9. Balancing eBooks and Physical Books Chemistry 12 Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 12 Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry 12 Guide
 - Setting Reading Goals Chemistry 12 Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry 12 Guide
 - Fact-Checking eBook Content of Chemistry 12 Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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