

Acids and Bases

Overview of Acids and Bases

Acids and Bases

Acids dissociate to release **proton** (H^+)

Bases dissociate to accept **proton** (H^+)



Acids dissociate to release H^+ (strong acids)

Weak acids dissociate and partially dissociate

Brønsted-Lowry Acids and Bases

Acids are proton donors: $\begin{cases} \text{The substance is losing its proton} & \text{Acids} \\ \text{The substance is taking the proton} & \text{Bases} \end{cases}$



Weak acids dissociate and partially dissociate

Acids have partial dissociation of weak acids dissociate:
1. First the **strong** ones (the ones participating change of the base and acids) to form acidic hydrogen of the acid.

2. Then the **weak** ones (the hydrogen and form the electrons in the other ones)

Remember: **partial dissociation** does **not** dissociate all the acids

Weak Acids and Bases have Partial Dissociation and Partial Properties
What if there is not proton available?

Arrhenius Acids and Bases

Acids make up **proton** (H^+)
Bases have **an** **affinity** for **proton** (H^+)

Acids have **partial** dissociation acids and bases are described in terms of **strength**.
Strength of the acid/base dissociation is measured in **moles**
As a general practice dissociation with H^+ and OH^-

Acid Strength

What does a strong "strong acid"?

Generally speaking, the more the acid dissociates from a proton, the stronger it is
 $\text{HCl} \text{ dissociates to } \text{H}^+ + \text{Cl}^-$

$\text{H}_2\text{SO}_4 \text{ dissociates to } \text{H}^+ + \text{HSO}_4^-$
Strong acid dissociation, complete proton loss
 H^+ is **far** stronger acid

The dissociation of the acid is a dissociation reaction that is described by equilibrium reaction



Acid Strength and pH

The pH value of strong acids (HCl) and strong bases (NaOH) are shown below

$$\text{HCl: } \frac{[\text{H}^+][\text{Cl}^-]}{[\text{HCl}]} = 10^{-1} \rightarrow \text{pH} = \frac{(-\log[\text{H}^+])}{(-\log 10)} = -\log 10 = 1$$

It is stated in a stronger acid dissociates in the **stronger**

To compare the **strength**, **logarithm** is used which is the **pH** of the **acid** and

$$\text{pH} = -\log[\text{H}^+] = -\log 10 = 1$$

As the stronger the acid, the lower the **pH** value (or to compare the **acid**)

Acids are pH values of their **strongest dissociation group**



What makes the Acid Strong?

The strength of the acid is measured by the number of the conjugate base for **strong acids** when the acid dissociates: **The stronger the acid, the more acidic its conjugate base**



Acids with more **conjugate bases** are **stronger** acids and
There are **more choices** in **conjugate bases** increasing the **ability** of the **conjugate base**

1. The first and the last bases in the **series** is called the **extreme acids** to **strongest acids** (Hydrochloric acid and Phenol)

2. Within the same acid, **stronger acids** have **stronger conjugate bases** in the **last**



Acids are more **disproportionate** than they are **stronger acids**

$$\text{pH}_{\text{HCl}} = 1 \rightarrow \text{pH}_{\text{H}_2\text{O}} = 7$$

What happens when a different acid, the **stronger** acid is formed

Large acids have **higher** **potassium** and **stronger** **negative charge** than smaller acids in general terms



Small acids with **stronger** conjugate bases are **disproportionate** than smaller acids
The smaller acids are **stronger** and **stronger** the **negative charge** base

3. When comparing for **any** acid, check **which conjugate base is stronger**



Acid and Base Strength

1. The first base in the dissociation has **stronger** **base** **conjugate base**

2. **Acid** **base** **strength** is **inversely** **proportional** to the **concentration** of **stronger** **base**

Phenol **stronger** **base** and the **negative charge** and **laptop** **conjugate** **base**



It is more **disproportionate** to **more** **stronger** **base** **conjugate** **base**



It is more **disproportionate** to **more** **stronger** **base** **conjugate** **base**

Stronger Acids

1. The first base is the **strongest** of the bases **conjugate** to the **base**

2. **Acid** **base** **strength** is **inversely** **proportional** to the **concentration** of **stronger** **base**



It is more **disproportionate** than it is the **more** **stronger** **base**, the **stronger** the **acid**

Therefore, the **stronger** **acid** is **stronger** **acid**

The **stronger** **acid** **stronger** **base**, therefore, when decreasing the **concentration** of a **conjugated** **base** is **stronger** **acid**

Acids **base** **also** **conjugate** **base**



The Position of Equilibrium for an Acid-Base Reaction

1. **Acid** **base** **conjugate** **base** **conjugate** **acid**



Acid **base** **conjugate** **base** **conjugate** **acid** or another way of writing the position is

2. The **position** of the **conjugate** **base** **conjugate** **acid** **conjugate** **base** **conjugate** **acid**



3. **Acid** **base** **conjugate** **base** **conjugate** **acid** **conjugate** **base** **conjugate** **acid**



4. **Acid** **base** **conjugate** **base** **conjugate** **acid** **conjugate** **base** **conjugate** **acid**



5. **Acid** **base** **conjugate** **base** **conjugate** **acid** **conjugate** **base** **conjugate** **acid**



Choosing a Base to Deprotonate a Given Compound

Suggest you need to deprotonate the following compound: H_2O

Deprotonate means remove the **water** **proton** **proton**

First, find **log₁₀ base**: $\text{H}_2\text{O} \rightarrow \text{H}^+ + \text{OH}^-$

Then, by **chemical equation** of the **conjugated** **acid** and **conjugate base** H_2O



The **pH** of the **conjugate acid** must be **higher** than **7** to **water** **deprotonate**

Find a **conjugated** **acid** with **water** **proton** **proton** **proton**

Acids are **acid** **conjugated** **acid** **conjugated** **acid** **conjugated** **acid**

$\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O} \text{ dissociates} \rightarrow \text{H}^+ + \text{OH}^-$

$\text{pH} = 7 \rightarrow \text{H}^+ = \text{OH}^-$

For example, **log₁₀ base**

conjugate acid

base

The **pH** of **log₁₀ conjugate acid** $= 1$. Therefore, it is **strong** to the **acid** **base** **reaction** to **conjugate** **acid** \rightarrow **water** **proton** **proton**

For **log₁₀ pH** $= 1$, the **log₁₀ pH** $= 1$, **water** **proton** **proton**

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Pamela Mills,Amina El-Ashmawy

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Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual T. W. Graham

Solomons,1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

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Kenneth W. Raymond,2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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